

**ORGANOSOL LIQUID TONER INCLUDING AMPHIPATHIC COPOLYMERIC  
BINDER HAVING CROSSLINKABLE FUNCTIONALITY**

**CROSS-REFERENCE TO RELATED APPLICATIONS**

This non-provisional Application claims the benefit of commonly assigned U.S. Provisional Application having serial number 60/437,881, filed on January 3, 2003, and titled **ORGANOSOL LIQUID TONER INCLUDING AMPHIPATHIC COPOLYMERIC BINDER HAVING CROSSLINKABLE FUNCTIONALITY**, which Application is incorporated herein by reference in its entirety.

**FIELD OF THE INVENTION**

The present invention relates to liquid toner compositions having utility in electrography. More particularly, the invention relates liquid electrographic liquid toners derived from organosols incorporating amphipathic copolymeric binder particles that include crosslinkable functionality.

**BACKGROUND OF THE INVENTION**

In electrographic and electrostatic printing processes (collectively electrographic processes), an electrostatic image is formed on the surface of a photoreceptive element or dielectric element, respectively. The photoreceptive element or dielectric element may be an intermediate transfer drum or belt or the substrate for the final toned image itself, as described by Schmidt, S. P. and Larson, J. R. in Handbook of Imaging Materials, Diamond, A. S., Ed: Marcel Dekker: New York; Chapter 6, pp 227-252, and U.S. Pat. Nos. 4,728,983, 4,321,404, and 4,268,598.

In electrostatic printing, a latent image is typically formed by (1) placing a charge image onto a dielectric element (typically the receiving substrate) in selected areas of the element with an electrostatic writing stylus or its equivalent to form a charge image, (2) applying toner to the charge image, and (3) fixing the toned image. An example of this type of process is described in U.S. Pat. No. 5,262,259.

In electrophotographic printing, also referred to as xerography, electrophotographic technology is used to produce images on a final image receptor, such as paper, film, or the like. Electrophotographic technology is incorporated into a wide range of equipment including photocopiers, laser printers, facsimile machines, and the like.

Electrophotography typically involves the use of a reusable, light sensitive, temporary image receptor, known as a photoreceptor, in the process of producing an electrophotographic image on a final, permanent image receptor. A representative electrophotographic process involves a series of steps to produce an image on a receptor, including charging, exposure, development, transfer, fusing, and cleaning, and erasure.

In the charging step, a photoreceptor is covered with charge of a desired polarity, either negative or positive, typically with a corona or charging roller. In the exposure step, an optical system, typically a laser scanner or diode array, forms a latent image by selectively discharging the charged surface of the photoreceptor in an imagewise manner corresponding to the desired image to be formed on the final image receptor. In the development step, toner particles of the appropriate polarity are generally brought into contact with the latent image on the photoreceptor, typically using a developer electrically-biased to a potential opposite in polarity to the toner polarity. The toner particles migrate to the photoreceptor and selectively adhere to the latent image via electrostatic forces, forming a toned image on the photoreceptor.

In the transfer step, the toned image is transferred from the photoreceptor to the desired final image receptor; an intermediate transfer element is sometimes used to effect transfer of the toned image from the photoreceptor with subsequent transfer of the toned image to a final image receptor. In the fusing step, the toned image on the final image receptor is heated to soften or melt the toner particles, thereby fusing the toned image to the final receptor. An alternative fusing method involves fixing the toner to the final receptor under high pressure with or without heat. In the cleaning step, residual toner remaining on the photoreceptor is removed.

Finally, in the erasing step, the photoreceptor charge is reduced to a substantially uniformly low value by exposure to light of a particular wavelength band, thereby

removing remnants of the original latent image and preparing the photoreceptor for the next imaging cycle.

Two types of toner are in widespread, commercial use: liquid toner and dry toner. The term “dry” does not mean that the dry toner is totally free of any liquid constituents, but connotes that the toner particles do not contain any significant amount of solvent, e.g., typically less than 10 weight percent solvent (generally, dry toner is as dry as is reasonably practical in terms of solvent content), and are capable of carrying a triboelectric charge. This distinguishes dry toner particles from liquid toner particles.

A typical liquid toner composition generally includes toner particles suspended or dispersed in a liquid carrier. The liquid carrier is typically a nonconductive dispersant, to avoid discharging the latent electrostatic image. Liquid toner particles are generally solvated to some degree in the liquid carrier (or carrier liquid), typically in more than 50 weight percent of a low polarity, low dielectric constant, substantially nonaqueous carrier solvent. Liquid toner particles are also typically smaller than dry toner particles. Because of their small particle size, ranging from about 5 microns to sub-micron, liquid toners are capable of producing very high-resolution toned images.

A typical toner particle for a liquid toner composition generally comprises a copolymeric binder and optionally one or more visual enhancement additives (for example, a colored pigment particle). The polymeric binder fulfills functions both during and after the electrophotographic process. With respect to processability, the character of the binder impacts charging and charge stability, flow, and fusing characteristics of the toner particles. These characteristics are important to achieve good performance during development, transfer, and fusing. After an image is formed on the final receptor, the nature of the binder (e.g. glass transition temperature, melt viscosity, molecular weight) and the fusing conditions (e.g. temperature, pressure and fuser configuration) impact durability (e.g. blocking and erasure resistance), adhesion to the receptor, gloss, and the like.

Polymeric binder materials suitable for use in liquid toner particles typically exhibit glass transition temperatures of about -24°C to 55°C, which is lower than the range of glass transition temperatures (50-100°C) typical for polymeric binders used in dry toner particles. In particular, some liquid toners are known to incorporate polymeric

binders exhibiting glass transition temperatures ( $T_g$ ) below room temperature (25°C) in order to rapidly self fix, e.g. by film formation, in the liquid electrophotographic imaging process; see e.g. U.S. 6,255,363. However, such liquid toners are also known to exhibit inferior image durability resulting from the low  $T_g$  (e.g. poor blocking and erasure resistance) after fusing the toned image to a final image receptor.

To overcome these durability deficiencies, polymeric materials selected for use in dry toners more typically exhibit a range of  $T_g$  of at least about 55-65°C in order to obtain good blocking resistance after fusing, yet typically require high fusing temperatures of about 200-250°C in order to soften or melt the toner particles and thereby adequately fuse the toner to the final image receptor. High fusing temperatures are a disadvantage for dry toners because of the long warm-up time and higher energy consumption associated with high temperature fusing and because of the risk of fire associated with fusing toner to paper at temperatures approaching the autoignition temperature of paper (233°C).

Although some liquid toners are known to use higher  $T_g$  (greater than or equal to about 60°C) polymeric binders, such toners are known to exhibit other problems related to the choice of polymeric binder, including image defects due to the inability of the liquid toner to rapidly self fix in the imaging process, poor charging and charge stability, poor stability with respect to agglomeration or aggregation in storage, poor sedimentation stability in storage, and the requirement that high fusing temperatures of about 200-250°C be used in order to soften or melt the toner particles and thereby adequately fuse the toner to the final image receptor.

In addition, some liquid and dry toners using high  $T_g$  polymeric binders are known to exhibit undesirable partial transfer (offset) of the toned image from the final image receptor to the fuser surface at temperatures above or below the optimal fusing temperature, requiring the use of low surface energy materials in the fuser surface or the application of fuser oils to prevent offset. Alternatively, various lubricants or waxes have been physically blended into the dry toner particles during fabrication to act as release or slip agents; however, because these waxes are not chemically bonded to the polymeric binder, they may adversely affect triboelectric charging of the toner particle or may migrate from the toner particle and contaminate the photoreceptor, an intermediate

transfer element, the fuser element, or other surfaces critical to the electrophotographic process.

In addition to the polymeric binder and the optional visual enhancement additive, liquid toner compositions can optionally include other additives. For example, charge control agents can be added to impart an electrostatic charge on the toner particles. Dispersing agents can be added to provide colloidal stability, aid fixing of the image, and provide charged or charging sites for the particle surface. Dispersing agents are commonly added to liquid toner compositions because toner particle concentrations are high (inter-particle distances are small) and electrical double-layer effects alone will not adequately stabilize the dispersion with respect to aggregation or agglomeration. Release agents can also be used to help prevent the toner from sticking to fuser rolls when those are used. Other additives include antioxidants, ultraviolet stabilizers, fungicides, bactericides, flow control agents, and the like.

One fabrication technique involves synthesizing an amphipathic copolymeric binder dispersed in a liquid carrier to form an organosol, then mixing the formed organosol with other ingredients to form a liquid toner composition. Typically, organosols are synthesized by nonaqueous dispersion polymerization of polymerizable compounds (e.g. monomers) to form copolymeric binder particles that are dispersed in a low dielectric hydrocarbon solvent (carrier liquid). These dispersed copolymer particles are sterically-stabilized with respect to aggregation by chemical bonding of a steric stabilizer (e.g. graft stabilizer), solvated by the carrier liquid, to the dispersed core particles as they are formed in the polymerization. Details of the mechanism of such steric stabilization are described in Napper, D.H., "Polymeric Stabilization of Colloidal Dispersions," Academic Press, New York, N.Y., 1983. Procedures for synthesizing self-stable organosols are described in "Dispersion Polymerization in Organic Media," K.E.J. Barrett, ed., John Wiley: New York, N.Y., 1975. Liquid toner compositions have been manufactured using dispersion polymerization in low polarity, low dielectric constant carrier solvents for use in making relatively low glass transition temperature ( $T_g \leq 30^\circ\text{C}$ ) film-forming liquid toners that undergo rapid self-fixing in the electrophotographic imaging process. See, e.g., U.S. Pat. No. 5,886,067 and 6,103,781. Organosols have also been prepared for use in making intermediate glass transition temperature ( $T_g$  between

30-55°C) liquid electrostatic toners for use in electrostatic stylus printers. See e.g. U.S. Pat. No. 6,255,363 B1. A representative non-aqueous dispersion polymerization method for forming an organosol is a free radical polymerization carried out when one or more ethylenically-unsaturated monomers, soluble in a hydrocarbon medium, are polymerized  
5 in the presence of a preformed, polymerizable solution polymer (e.g. a graft stabilizer or “living” polymer). See U.S. Pat. No. 6,255,363.

Once the organosol has been formed, one or more additives can be incorporated, as desired. For example, one or more visual enhancement additives and/or charge control agents can be incorporated. The composition can then subjected to one or more mixing  
10 processes, such as homogenization, microfluidization, ball-milling, attritor milling, high energy bead (sand) milling, basket milling or other techniques known in the art to reduce particle size in a dispersion. The mixing process acts to break down aggregated visual enhancement additive particles, when present, into primary particles (having a diameter in the range of 0.05 to 1.0 microns) and may also partially shred the dispersed  
15 copolymeric binder into fragments that can associate with the surface of the visual enhancement additive.

According to this embodiment, the dispersed copolymer or fragments derived from the copolymer then associate with the visual enhancement additive, for example, by adsorbing to or adhering to the surface of the visual enhancement additive, thereby  
20 forming toner particles. The result is a sterically-stabilized, nonaqueous dispersion of toner particles having a size in the range of about 0.1 to 2.0 microns, with typical toner particle diameters in the range 0.1 to 0.5 microns. In some embodiments, one or more charge control agents can be added after mixing, if desired.

Several characteristics of liquid toner compositions are important to provide high  
25 quality images. Toner particle size and charge characteristics are especially important to form high quality images with good resolution. Further, rapid self-fixing of the toner particles is an important requirement for some liquid electrophotographic printing applications, e.g. to avoid printing defects (such as smearing or trailing-edge tailing) and incomplete transfer in high-speed printing. Another important consideration in  
30 formulating a liquid toner composition relates to the durability and archivability of the image on the final receptor. Erasure resistance, e.g. resistance to removal or damage of

the toned image by abrasion, particularly by abrasion from natural or synthetic rubber erasers commonly used to remove extraneous pencil or pen markings, is a desirable characteristic of liquid toner particles.

Resistance of the image on the final image receptor to damage by blocking to the  
5 receptor (or to other toned surfaces) is another desirable characteristic of liquid toner particles. Therefore, another important consideration in formulating a liquid toner is the tack of the image on the final receptor. It is desirable for the image on the final receptor material to be essentially tack-free over a fairly wide range of temperatures. If the image has a residual tack, then the image can become embossed or picked off when placed in  
10 contact with another surface (also referred to as blocking). This is particularly a problem when printed sheets are placed in a stack.

To address this concern, a film laminate or protective layer may be placed over the surface of the image. This laminate often acts to increase the effective dot gain of the image, thereby interfering with the color rendition of a color composite. In addition,  
15 lamination of a protective layer over a final image surface adds both extra cost of materials and extra process steps to apply the protective layer, and may be unacceptable for certain printing applications (e.g. plain paper copying or printing).

Another method to improve the durability of liquid toned images and address the drawbacks of lamination is described in U.S. Patent No. 6,103,781. U.S. Patent No.  
20 6,103,781 describes a liquid ink composition containing organosols having side-chain or main-chain crystallizable polymeric moieties. At column 6, lines 53 - 60, the authors describe a binder resin that is an amphipathic copolymer dispersed in a liquid carrier (also known as an organosol) that includes a high molecular weight (co)polymeric steric stabilizer covalently bonded to an insoluble, thermoplastic (co)polymeric core. The steric  
25 stabilizer includes a crystallizable polymeric moiety that is capable of independently and reversibly crystallizing at or above room temperature (22°C).

According to the authors, superior stability of the dispersed toner particles with respect to aggregation is obtained when at least one of the polymers or copolymers (denoted as the stabilizer) is an amphipathic substance containing at least one oligomeric  
30 or polymeric component having a weight-average molecular weight of at least 5,000 which is solvated by the liquid carrier. In other words, the selected stabilizer, if present

as an independent molecule, would have some finite solubility in the liquid carrier. Generally, this requirement is met if the absolute difference in Hildebrand solubility parameter between the steric stabilizer and the solvent is less than or equal to  $3.0 \text{ MPa}^{1/2}$ .

As described in U.S. Patent No. 6,103,781, the composition of the insoluble resin core is preferentially manipulated such that the organosol exhibits an effective glass transition temperature ( $T_g$ ) of less than  $22^\circ\text{C}$ , more preferably less than  $6^\circ\text{C}$ . Controlling the glass transition temperature allows one to formulate an ink composition containing the resin as a major component to undergo rapid film formation (rapid self-fixing) in liquid electrophotographic printing or imaging processes using offset transfer processes carried out at temperatures greater than the core  $T_g$ , preferably at or above  $22^\circ\text{C}$ . (Column 10, lines 36-46).

## SUMMARY OF THE INVENTION

The present invention relates to liquid toner compositions having utility in electrographic applications. In particular, the present invention relates to organosol liquid toner compositions comprising binder particles dispersed in a nonaqueous liquid carrier, wherein the particles are derived from ingredients comprising one or more crosslinkable amphipathic copolymer(s). The organosol is easily combined with additional ingredients, such as one or more visual enhancement additives and other desired ingredients, and subjected to mixing processes to form a liquid toner composition.

The compositions provide beneficial performance characteristics attributable to both low and high  $T_g$  liquid toner formulations, which benefits conventionally have been mutually exclusive in many regards. Prior to being crosslinked, some embodiments of the copolymers of the present invention can have lower  $T_g$  characteristics allowing formulating at higher solids content, enhanced self-fixing, higher resolution imaging, faster drying, lower fusing temperatures, and the like. Such performance advantages are generally not as readily available when using higher  $T_g$  materials. After being crosslinked, e.g., at some point after image development, the resultant images offer good durability, anchoring via crosslinking to substrates, other image layers, coverlays or the like, blocking resistance, and temperature resistance. Such performance advantages are



generally not as readily available when using lower Tg materials lacking crosslinking functionality.

As used herein, the term “amphipathic” refers to a copolymer having a combination of portions having distinct solubility and dispersibility characteristics in a desired liquid carrier that is used to make the copolymer and/or used in the course of preparing the liquid toner particles. Preferably, the liquid carrier is selected such that at least one portion (also referred to herein as S material or portion(s)) of the copolymer is more solvated by the carrier while at least one other portion (also referred to herein as D material or portion(s)) of the copolymer constitutes more of a dispersed phase in the carrier.

In preferred embodiments, the amphipathic copolymer is polymerized in situ in the desired liquid carrier as this yields substantially monodisperse copolymeric particles suitable for use in liquid toner compositions with little, if any, need for subsequent comminuting or classifying. The resulting organosol is then converted into toner particles by mixing the organosol with other optional ingredients, such as at least one visual enhancement additive and other desired ingredients. During such combination, ingredients comprising the visual enhancement particles and the amphipathic copolymer will tend to self-assemble into composite toner particles. Specifically, it is believed that the D portion of the copolymer will tend to physically and/or chemically interact with the surface of the visual enhancement additive, while the S portion helps promote dispersion in the carrier without use of a separate surfactant or dispersant.

Additionally, a wide range of liquid carrier soluble or dispersible monomers may be used to form the organosol by a variety of substantially nonaqueous polymerization methods. Preferably, substantially nonaqueous dispersion polymerization is used to polymerize monomers using free radical polymerization methods as desired. As used herein, “substantially nonaqueous polymerization methods” refers to polymerization methods in an organic solvent containing at most a minor portion of water.

In one aspect, the present invention relates to a liquid electrographic toner composition comprising a liquid carrier having a Kauri-Butanol number less than 30. A plurality of toner particles is dispersed in the liquid carrier. The toner particles comprise at least one amphipathic copolymer comprising one or more S material portions and one

or more D material portions. The toner particles comprise complementary crosslinkable functionalities, which may be the same or different, wherein at least a portion of the crosslinkable functionality is incorporated into the amphipathic copolymer.

In another aspect, the present invention relates to a liquid electrographic toner composition comprising a liquid carrier having a Kauri-Butanol number less than 30. A first plurality of toner particles dispersed in the liquid carrier, wherein the first plurality of toner particles comprise a first amphipathic copolymer comprising one or more S material portions and one or more D material portions, and wherein the first amphipathic copolymer comprises a first crosslinkable functionality. A second plurality of toner particles dispersed in the liquid carrier, wherein the second plurality of toner particles comprises a second amphipathic copolymer comprising one or more S material portions and one or more D material portions. The second amphipathic copolymer comprises a second crosslinkable functionality, wherein the first and second crosslinkable functionalities are complementary.

In another aspect, the present invention relates to a method of making a liquid electrographic toner composition. An organosol is provided that comprises a plurality of toner particles dispersed in a liquid carrier, wherein the toner particles comprise at least one amphipathic copolymer. The amphipathic copolymer comprises one or more S material portions and one or more D material portions. The amphipathic copolymer also comprises crosslinkable functionality. The organosol is mixed with one or more additives under conditions effective to form a dispersion.

In another aspect, the present invention relates to a method of electrographically forming an image on a substrate surface. liquid toner composition is provided, wherein the liquid toner composition comprises an organosol. The organosol comprises a plurality of toner particles dispersed in a liquid carrier, wherein the toner particles comprise at least one amphipathic copolymer comprising one or more S material portions and one or more D material portions. The amphipathic copolymer comprises crosslinkable functionality. An image comprising the toner particles is caused to be formed on the substrate surface. The amphipathic copolymer is crosslinked.

## BRIEF DESCRIPTION OF THE DRAWINGS

Fig. 1a schematically shows one embodiment of an organosol of the present invention comprising a crosslinkable, amphipathic copolymer.

5 Fig. 1b schematically shows one embodiment of an organosol of the present invention comprising a crosslinkable, amphipathic copolymer.

Fig. 1c schematically shows one embodiment of an organosol of the present invention comprising a crosslinkable, amphipathic copolymer.

10 Fig. 1d schematically shows one embodiment of an organosol of the present invention comprising a crosslinkable, amphipathic copolymer.

Fig. 2a schematically shows one embodiment of an organosol of the present invention comprising a combination of crosslinkable, amphipathic copolymers.

Fig. 2b schematically shows one embodiment of an organosol of the present invention comprising a combination of crosslinkable, amphipathic copolymers.

15 Fig. 3 schematically shows a device comprising a tamper-resistant image formed using ingredients comprising a liquid toner of the present invention.

Fig. 4 is a graph showing erasure resistance v. Crock cloth passes for the data obtained in Examples 14, 15 and 19.

## 20 DETAILED DESCRIPTION OF PRESENTLY PREFERRED EMBODIMENTS

The embodiments of the present invention described below are not intended to be exhaustive or to limit the invention to the precise forms disclosed in the following detailed description. Rather, the embodiments are chosen and described so that others  
25 skilled in the art can appreciate and understand the principles and practices of the present invention.

Organosol liquid toner compositions of the present invention generally comprise toner particles dispersed in a nonaqueous liquid carrier, wherein the particles are derived from ingredients comprising an amphipathic copolymer. Preferably, the nonaqueous  
30 liquid carrier of the organosol is selected such that at least one portion (also referred to herein as the S material or portion) of the amphipathic copolymer is more solvated by the carrier while at least one other portion (also referred to herein as the D material or

portion) of the copolymer constitutes more of a dispersed phase in the carrier. In other words, preferred copolymers of the present invention comprise S and D material having respective solubilities in the desired liquid carrier that are sufficiently different from each other such that the S blocks tend to be more solvated by the carrier while the D blocks  
5 tend to be more dispersed in the carrier. More preferably, the S blocks are soluble in the liquid carrier while the D blocks are insoluble. In particularly preferred embodiments, the D material phase separates from the liquid carrier, forming dispersed particles.

From one perspective, the polymer particles when dispersed in the liquid carrier may be viewed as having a core/shell structure in which the D material tends to be in the  
10 core, while the S material tends to be in the shell. The S material thus functions as a dispersing aid, steric stabilizer or graft copolymer stabilizer, to help stabilize dispersions of the copolymer particles in the liquid carrier. Consequently, the S material may also be referred to herein as a "graft stabilizer." The core/shell structure of the binder particles tends to be retained when the particles are dried when incorporated into dry toner  
15 particles.

The solubility of a material, or a portion of a material such as a copolymeric portion, may be qualitatively and quantitatively characterized in terms of its Hildebrand solubility parameter. The Hildebrand solubility parameter refers to a solubility parameter represented by the square root of the cohesive energy density of a material, having units  
20 of (pressure)<sup>1/2</sup>, and being equal to  $(\Delta H/RT)^{1/2} / V^{1/2}$ , where  $\Delta H$  is the molar vaporization enthalpy of the material, R is the universal gas constant, T is the absolute temperature, and V is the molar volume of the solvent. Hildebrand solubility parameters are tabulated for solvents in Barton, A. F. M., Handbook of Solubility and Other Cohesion Parameters, 2d Ed. CRC Press, Boca Raton, Fla., (1991), for monomers and representative polymers  
25 in Polymer Handbook, 3rd Ed., J. Brandrup & E. H. Immergut, Eds. John Wiley, N.Y., pp 519-557 (1989), and for many commercially available polymers in Barton, A. F. M., Handbook of Polymer-Liquid Interaction Parameters and Solubility Parameters, CRC Press, Boca Raton, Fla., (1990).

The degree of solubility of a material, or portion thereof, in a liquid carrier may  
30 be predicted from the absolute difference in Hildebrand solubility parameters between the material, or portion thereof, and the liquid carrier. A material, or portion thereof, will be

fully soluble or at least in a highly solvated state when the absolute difference in Hildebrand solubility parameter between the material, or portion thereof, and the liquid carrier is less than approximately  $1.5 \text{ MPa}^{1/2}$ . On the other hand, when the absolute difference between the Hildebrand solubility parameters exceeds approximately  $3.0 \text{ MPa}^{1/2}$ , the material, or portion thereof, will tend to phase separate from the liquid carrier, forming a dispersion. When the absolute difference in Hildebrand solubility parameters is between  $1.5 \text{ MPa}^{1/2}$  and  $3.0 \text{ MPa}^{1/2}$ , the material, or portion thereof, is considered to be weakly solvatable or marginally insoluble in the liquid carrier.

Consequently, in preferred embodiments, the absolute difference between the respective Hildebrand solubility parameters of the S portion(s) of the copolymer and the liquid carrier is less than  $3.0 \text{ MPa}^{1/2}$ , preferably less than about  $2.0 \text{ MPa}^{1/2}$ , more preferably less than about  $1.5 \text{ MPa}^{1/2}$ . Additionally, it is also preferred that the absolute difference between the respective Hildebrand solubility parameters of the D portion(s) of the copolymer and the liquid carrier is greater than  $2.3 \text{ MPa}^{1/2}$ , preferably greater than about  $2.5 \text{ MPa}^{1/2}$ , more preferably greater than about  $3.0 \text{ MPa}^{1/2}$ , with the proviso that the difference between the respective Hildebrand solubility parameters of the S and D portion(s) is at least about  $0.4 \text{ MPa}^{1/2}$ , more preferably at least about  $1.0 \text{ MPa}^{1/2}$ . Because the Hildebrand solubility of a material may vary with changes in temperature, such solubility parameters are preferably determined at a desired reference temperature such as at  $25^\circ\text{C}$ .

Those skilled in the art understand that the Hildebrand solubility parameter for a copolymer, or portion thereof, may be calculated using a volume fraction weighting of the individual Hildebrand solubility parameters for each monomer comprising the copolymer, or portion thereof, as described for binary copolymers in Barton A. F. M., Handbook of Solubility Parameters and Other Cohesion Parameters, CRC Press, Boca Raton, p 12 (1990). The magnitude of the Hildebrand solubility parameter for polymeric materials is also known to be weakly dependent upon the weight average molecular weight of the polymer, as noted in Barton, pp 446-448. Thus, there will be a preferred molecular weight range for a given polymer or portion thereof in order to achieve desired solvating or dispersing characteristics. Similarly, the Hildebrand solubility parameter for

a mixture may be calculated using a volume fraction weighting of the individual Hildebrand solubility parameters for each component of the mixture.

In addition, we have defined our invention in terms of the calculated solubility parameters of the monomers and solvents obtained using the group contribution method developed by Small, P. A., J. Appl. Chem., 3, 71 (1953) using Small's group contribution values listed in Table 2.2 on page VII/525 in the Polymer Handbook, 3rd Ed., J. Brandrup & E. H. Immergut, Eds. John Wiley, New York, (1989). We have chosen this method for defining our invention to avoid ambiguities which could result from using solubility parameter values obtained with different experimental methods. In addition, Small's group contribution values will generate solubility parameters that are consistent with data derived from measurements of the enthalpy of vaporization, and therefore are completely consistent with the defining expression for the Hildebrand solubility parameter. Since it is not practical to measure the heat of vaporization for polymers, monomers are a reasonable substitution.

For purposes of illustration, Table I lists Hildebrand solubility parameters for some common solvents used in an electrographic toner and the Hildebrand solubility parameters and glass transition temperatures (based on their high molecular weight homopolymers) for some common monomers used in synthesizing organosols.

**TABLE I**

Hildebrand Solubility Parameters Solvent Values at 25°C		
Solvent Name	Kauri-Butanol Number by ASTM Method D1133- 54T (ml)	Hildebrand Solubility Parameter (MPa <sup>1/2</sup> )
Norpar <sup>TM</sup> 15	18	13.99
Norpar <sup>TM</sup> 13	22	14.24
Norpar <sup>TM</sup> 12	23	14.30
Isopar <sup>TM</sup> V	25	14.42
Isopar <sup>TM</sup> G	28	14.60
Exxsol <sup>TM</sup> D80	28	14.60
Source: Calculated from equation #31 of Polymer Handbook, 3 <sup>rd</sup> Ed., J. Brandrup E. H. Immergut, Eds. John Wiley, NY, p. VII/522 (1989).		

Monomer Values at 25°C		
Monomer Name	Hildebrand Solubility Parameter (MPa <sup>1/2</sup> )	Glass Transition Temperature (°C)*
3,3,5-Trimethyl Cyclohexyl Methacrylate	16.73	125
Isobornyl Methacrylate	16.90	110
Isobornyl Acrylate	16.01	94
n-Behenyl acrylate	16.74	-65 (58 m.p.)**
n-Octadecyl Methacrylate	16.77	-100 (45 m.p.) **
n-Octadecyl Acrylate	16.82	-55
Lauryl Methacrylate	16.84	-65
Lauryl Acrylate	16.95	-30
2-Ethylhexyl Methacrylate	16.97	-10
2-Ethylhexyl Acrylate	17.03	-55
n-Hexyl Methacrylate	17.13	-5
t-Butyl Methacrylate	17.16	107
n-Butyl Methacrylate	17.22	20
n-Hexyl Acrylate	17.30	-60
n-Butyl Acrylate	17.45	-55
Ethyl Methacrylate	17.62	65
Ethyl Acrylate	18.04	-24
Methyl Methacrylate	18.17	105
Styrene	18.05	100
Calculated using Small's Group Contribution Method, Small, P.A. Journal of Applied Chemistry 3 p. 71 (1953). Using Group Contributions from Polymer Handbook, 3 <sup>rd</sup> Ed., J. Brandrup E. H. Immergut, Eds., John Wiley, NY, p. VII/525 (1989). *Polymer Handbook, 3 <sup>rd</sup> Ed., J. Brandrup E. H. Immergut, Eds., John Wiley, NY, pp. VII/209-277 (1989). The T <sub>g</sub> listed is for the homopolymer of the respective monomer. ** m.p. refers to melting point for selected Polymerizable Crystallizable Compounds.		

The liquid carrier is a substantially nonaqueous solvent or solvent blend. In other words, only a minor component (generally less than 25 weight percent) of the liquid carrier comprises water. Preferably, the substantially nonaqueous liquid carrier comprises less than 20 weight percent water, more preferably less than 10 weight percent water, even more preferably less than 3 weight percent water, most preferably less than one weight percent water.

The substantially nonaqueous carrier liquid may be selected from a wide variety of materials, or combination of materials, which are known in the art, but preferably has a Kauri-butanol number less than 30 ml. The liquid is preferably oleophilic, chemically stable under a variety of conditions, and electrically insulating. Electrically insulating refers to a liquid carrier having a low dielectric constant and a high electrical resistivity.

Preferably, the liquid carrier has a dielectric constant of less than 5; more preferably less than 3. Electrical resistivities of carrier liquids are typically greater than  $10^9$  Ohm-cm; more preferably greater than  $10^{10}$  Ohm-cm. In addition, the liquid carrier desirably is chemically inert in most embodiments with respect to the ingredients used to formulate the toner particles.

Examples of suitable liquid carriers include aliphatic hydrocarbons (n-pentane, hexane, heptane and the like), cycloaliphatic hydrocarbons (cyclopentane, cyclohexane and the like), aromatic hydrocarbons (benzene, toluene, xylene and the like), halogenated hydrocarbon solvents (chlorinated alkanes, fluorinated alkanes, chlorofluorocarbons and the like) silicone oils and blends of these solvents. Preferred carrier liquids include branched paraffinic solvent blends such as Isopar™ G, Isopar™ H, Isopar™ K, Isopar™ L, Isopar™ M and Isopar™ V (available from Exxon Corporation, NJ), and most preferred carriers are the aliphatic hydrocarbon solvent blends such as Norpar™ 12, Norpar™ 13 and Norpar™ 15 (available from Exxon Corporation, NJ).

As used herein, the term “copolymer” encompasses both oligomeric and polymeric materials, and encompasses copolymers incorporating two or more monomers. As used herein, the term “monomer” means a relatively low molecular weight material (i.e., generally having a molecular weight less than about 500 Daltons) having one or more polymerizable groups. “Oligomer” means a relatively intermediate sized molecule incorporating two or more monomers and generally having a molecular weight of from about 500 up to about 10,000 Daltons. “Polymer” means a relatively large material comprising a substructure formed two or more monomeric, oligomeric, and/or polymeric constituents and generally having a molecular weight greater than about 10,000 Daltons.

The term “macromer” or “macromonomer” refers to an oligomer or polymer having a terminal polymerizable moiety. “Polymerizable crystallizable compound” or “PCC” refers to compounds capable of undergoing polymerization to produce a copolymer wherein at least a portion of the copolymer is capable of undergoing reversible crystallization over a reproducible and well-defined temperature range (e.g. the copolymer exhibits a melting and freezing point as determined, for example, by differential scanning calorimetry). PCC’s may include monomers, functional oligomers, functional pre-polymers, macromers or other compounds able to undergo polymerization



to form a copolymer. The term “molecular weight” as used throughout this specification means weight average molecular weight unless expressly noted otherwise.

The weight average molecular weight of the amphipathic copolymer of the present invention may vary over a wide range, and may impact imaging performance.

5 The polydispersity of the copolymer also may impact imaging and transfer performance of the resultant dry toner material. Because of the difficulty of measuring molecular weight for an amphipathic copolymer, the particle size of the dispersed copolymer (organosol) may instead be correlated to imaging and transfer performance of the resultant dry toner material. Generally, the volume mean particle diameter ( $D_v$ ) of the  
10 dispersed graft copolymer particles, determined by laser diffraction particle size measurement, should be in the range 0.1-100 microns, more preferably 0.5-50 microns, even more preferably 1.0-20 microns, and most preferably 2-10 microns.

In addition, a correlation exists between the molecular weight of the solvatable or soluble S portion of the graft copolymer, and the imaging and transfer performance of the  
15 resultant toner. Generally, the S portion of the copolymer has a weight average molecular weight in the range of 1000 to about 1,000,000 Daltons, preferably 5000 to 400,000 Daltons, more preferably 50,000 to 300,000 Daltons. It is also generally desirable to maintain the polydispersity (the ratio of the weight-average molecular weight to the number average molecular weight) of the S portion of the copolymer below 15,  
20 more preferably below 5, most preferably below 2.5. It is a distinct advantage of the present invention that copolymer particles with such lower polydispersity characteristics for the S portion are easily made in accordance with the practices described herein, particularly those embodiments in which the copolymer is formed in the liquid carrier in situ.

25 The relative amounts of S and D portions in a copolymer can impact the solvating and dispersability characteristics of these portions. For instance, if too little of the S portion(s) are present, the copolymer may have too little stabilizing effect to sterically-stabilize the organosol with respect to aggregation as might be desired. If too little of the D portion(s) are present, the small amount of D material may be too soluble in the liquid  
30 carrier such that there may be insufficient driving force to form a distinct particulate, dispersed phase in the liquid carrier. The presence of both a solvated and dispersed phase

helps the ingredients of particles self assemble in situ with exceptional uniformity among separate particles. Balancing these concerns, the preferred weight ratio of D material to S material is in the range of 1:20 to 20:1, preferably 1:1 to 15:1, more preferably 2:1 to 10:1, and most preferably 4:1 to 8:1.

5           Glass transition temperature,  $T_g$ , refers to the temperature at which a (co)polymer, or portion thereof, changes from a hard, glassy material to a rubbery, or viscous, material, corresponding to a dramatic increase in free volume as the (co)polymer is heated. The  $T_g$  can be calculated for a (co)polymer, or portion thereof, using known  $T_g$  values for the high molecular weight homopolymers (see, e.g., Table I herein) and the Fox equation  
10       expressed below:

$$1/T_g = w_1/T_{g1} + w_2/T_{g2} + \dots w_i/T_{gi}$$

wherein each  $w_n$  is the weight fraction of monomer "n" and each  $T_{gn}$  is the absolute glass  
15       transition temperature (in degrees Kelvin) of the high molecular weight homopolymer of monomer "n" as described in Wicks, A. W., F. N. Jones & S. P. Pappas, Organic Coatings 1, John Wiley, NY, pp 54-55 (1992).

In the practice of the present invention, values of  $T_g$  for the D or S portion of the copolymer were determined using the Fox equation above, although the  $T_g$  of the  
20       copolymer as a whole may be determined experimentally using, for example, differential scanning calorimetry. The glass transition temperatures ( $T_g$ 's) of the S and D portions may vary over a wide range and may be independently selected to enhance manufacturability and/or performance of the resulting dry toner particles. The  $T_g$ 's of the S and D portions will depend to a large degree upon the type of monomers constituting  
25       such portions. Consequently, to provide a copolymer material with higher  $T_g$ , one can select one or more higher  $T_g$  monomers with the appropriate solubility characteristics for the type of copolymer portion (D or S) in which the monomer(s) will be used. Conversely, to provide a copolymer material with lower  $T_g$ , one can select one or more lower  $T_g$  monomers with the appropriate solubility characteristics for the type of portion  
30       in which the monomer(s) will be used.

For copolymers in which the D portion comprises a major portion of the copolymer, the  $T_g$  of the D portion will dominate the  $T_g$  of the copolymer as a whole. For such copolymers useful in liquid toner applications, it is preferred that the  $T_g$  of the D portion fall in the range of  $-25^{\circ}\text{C}$  to  $105^{\circ}\text{C}$ , more preferably  $0^{\circ}\text{C}$  to  $85^{\circ}\text{C}$ , most preferably  $8^{\circ}$  to  $65^{\circ}\text{C}$ . Use of low  $T_g$  D material is desirable to enhance properties such as drying performance, higher solids content in the liquid toner, self-fixing, reduced fusing temperatures, and the like. However, notwithstanding such benefits, using D material with a  $T_g$  that is too low can cause performance issues either with respect to blocking resistance, erasure resistance, or the like. It is a distinct advantage of the present invention that crosslinkable D material having low  $T_g$  characteristics, e.g., a  $T_g$  below about  $50^{\circ}\text{C}$ , more preferably below about  $30^{\circ}\text{C}$  may be used in liquid toner. Once an image is formed using a liquid toner of the present invention comprising low  $T_g$ , crosslinkable D material, the D material can be crosslinked, resulting in an image that is durable, temperature resistance, and highly resistant to blocking. In practical effect, the present invention allows the benefits of both low  $T_g$  and high  $T_g$  D material to be achieved from the same liquid toner formulation.

The S material most typically has relatively low  $T_g$  characteristics, as many of the monomers useful for forming S material are low  $T_g$  monomers. However, blocking with respect to the S portion material is not as significant an issue inasmuch as preferred copolymers comprise a majority of the D portion material. Consequently, the  $T_g$  of the D portion material will dominate the effective  $T_g$  of the copolymer as a whole. Additionally, S material of the present invention may be crosslinkable, so that blocking problems associated with the uncured S material are readily alleviated by crosslinking. However, if the  $T_g$  of the S portion is too low, then the particles might tend to aggregate. On the other hand, if the  $T_g$  is too high, then the requisite fusing temperature may be too high. Balancing these concerns, the S portion material is preferably formulated to have a  $T_g$  in the range from at least  $-65^{\circ}\text{C}$  to about  $60^{\circ}\text{C}$ , preferably at least  $-10^{\circ}\text{C}$  to about  $50^{\circ}\text{C}$ , more preferably at least  $0^{\circ}\text{C}$  to about  $50^{\circ}\text{C}$ .

It is understood that the desired performance criteria for the self-fixing characteristics of a liquid toner will depend to a great extent upon the nature of the imaging process. For example, rapid self-fixing of the toner to form a cohesive film may

not be required or even desired in an electrographic imaging process if the image is not subsequently transferred to a final receptor, or if the transfer is effected by means (e.g. electrostatic transfer) not requiring a film formed toner on a temporary image receptor (e.g. a photoreceptor). Similarly, in multi-color (or multi-pass) electrostatic printing

5 wherein a stylus is used to generate a latent electrostatic image directly upon a dielectric receptor that serves as the final toner receptor material, a rapidly self-fixing toner film may be undesirably removed in passing under the stylus. This head scraping can be reduced or eliminated by manipulating the effective glass transition temperature of the organosol. For liquid electrographic (electrostatic) toners, particularly liquid toners  
10 developed for use in direct electrostatic printing processes, the D portion of the organosol is preferably provided with a sufficiently high  $T_g$  such that the resultant copolymer exhibits an effective glass transition temperature of from about 15°C to about 55°C.

A wide variety of one or more different monomeric, oligomeric and/or polymeric materials may be independently incorporated into the S and D portions, as desired.

15 Representative examples of suitable materials include free radically polymerized material (also referred to as vinyl copolymers or (meth) acrylic copolymers in some embodiments), polyurethanes, polyester, epoxy, polyamide, polyimide, polysiloxane, fluoropolymer, polysulfone, combinations of these, and the like. Preferred S and D  
20 portions are derived from free radically polymerizable material. In the practice of the present invention, “free radically polymerizable” refers to monomers, oligomers, and/or polymers having functionality directly or indirectly pendant from a monomer, oligomer, or polymer backbone (as the case may be) that participate in polymerization reactions via a free radical mechanism. Representative examples of such functionality includes  
25 (meth)acrylate groups, olefinic carbon-carbon double bonds, allyloxy groups, alpha-methyl styrene groups, (meth)acrylamide groups, cyanate ester groups, vinyl ether groups, combinations of these, and the like. The term “(meth)acryl”, as used herein, encompasses acryl and/or methacryl.

Free radically polymerizable monomers, oligomers, and/or polymers are advantageously used to form the copolymer in that so many different types are  
30 commercially available and may be selected with a wide variety of desired characteristics that help provide one or more desired performance characteristics. Free radically

polymerizable monomers, oligomers, and/or monomers suitable in the practice of the present invention may include one or more free radically polymerizable moieties.

Representative examples of monofunctional, free radically polymerizable monomers include styrene, alpha-methylstyrene, substituted styrene, vinyl esters, vinyl ethers, N-vinyl-2-pyrrolidone, (meth)acrylamide, vinyl naphthalene, alkylated vinyl naphthalenes, alkoxy vinyl naphthalenes, N-substituted (meth)acrylamide, octyl (meth)acrylate, nonylphenol ethoxylate (meth)acrylate, N-vinyl pyrrolidone, isononyl (meth)acrylate, isobornyl (meth)acrylate, 2-(2-ethoxyethoxy)ethyl (meth)acrylate, 2-ethylhexyl (meth)acrylate, beta-carboxyethyl (meth)acrylate, isobutyl (meth)acrylate, cycloaliphatic epoxide, alpha-epoxide, 2-hydroxyethyl (meth)acrylate, (meth)acrylonitrile, maleic anhydride, itaconic acid, isodecyl (meth)acrylate, lauryl (dodecyl) (meth)acrylate, stearyl (octadecyl) (meth)acrylate, behenyl (meth)acrylate, n-butyl (meth)acrylate, methyl (meth)acrylate, ethyl (meth)acrylate, hexyl (meth)acrylate, (meth)acrylic acid, N-vinylcaprolactam, stearyl (meth)acrylate, hydroxy functional caprolactone ester (meth)acrylate, isooctyl (meth)acrylate, hydroxyethyl (meth)acrylate, hydroxymethyl (meth)acrylate, hydroxypropyl (meth)acrylate, hydroxyisopropyl (meth)acrylate, hydroxybutyl (meth)acrylate, hydroxyisobutyl (meth)acrylate, tetrahydrofurfuryl (meth)acrylate, isobornyl (meth)acrylate, glycidyl (meth)acrylate vinyl acetate, combinations of these, and the like.

Preferred copolymers of the present invention may be formulated with one or more radiation curable monomers or combinations thereof that help the free radically polymerizable compositions and/or resultant cured compositions to satisfy one or more desirable performance criteria. Advantageously, for example, the amphipathic copolymer(s) of the present invention incorporate monomeric, oligomeric, and/or polymeric constituents that provide the resultant copolymers with crosslinkable functionality. The crosslinkable functionality allows the copolymers to be thermosetting, e.g., chemically crosslinkable.

The liquid toner composition(s) including the crosslinkable (thermosetting) amphipathic copolymers may be formulated so that crosslinking can be internal and/or external. As used herein, internally crosslinkable means that the crosslinkable functionality incorporated into the amphipathic copolymer chemically crosslinks with

complementary, crosslinkable functionality on the same copolymer, with or without an initiator, catalyst, and/or crosslinking agent. Externally crosslinkable means that the crosslinkable functionality on a particular amphipathic copolymer chemically crosslinks with complementary crosslinkable functionality on a different polymer material, which  
5 may or may not be other amphipathic copolymeric material, and which may or may not occur with one or more initiator(s), catalyst(s), and/or crosslinking agent(s). When such functionalized copolymers are incorporated into liquid toners and then used to form images, crosslinks are readily formed within the same image layer and/or among two or more image layers. For example, Fig. 3 (described further below) schematically  
10 illustrates an embodiment of the invention in which interlayer crosslinking is used to provide tamper-resistant images.

The crosslinkable functionality can include one or more kinds of pendant, chemically reactive moieties that chemically react with complementary, chemically reactive moieties upon being crosslinked. The complementary, chemically reactive  
15 moieties may be the same or different depending upon the nature of the moieties and the desired chemical linkage that forms as a consequence of crosslinking. Complementary moieties are those that chemically react (optionally in the presence of an initiator, catalysis, crosslinking agent, or the like) to form a variety of inter and/or intrapolymeric linkages such as urethane linkages, ester linkages, urea linkages, amide linkages, epoxy  
20 linkages, sulfone linkages, siloxane linkages, imide linkages, olefinic linkages, acrylic linkages, combinations of these and the like. Particularly preferred complementary, chemically reactive moieties include OH and NCO moieties, which crosslink to form urethane linkages, OH and carboxylic acid or acid salt moieties, which crosslink to form ester linkages, amine (either secondary or primary) and NCO moieties which crosslink to  
25 form amide linkages, amine and carboxylic acid or salt moieties which crosslink to form amide linkages, epoxy and amine (either secondary or primary) moieties that react together, combinations of these, and the like.

Particularly preferred complementary, chemically reactive moieties are those that chemically crosslink at the desired rate and to the desired degree only upon being  
30 subjected to a particular crosslinking event. Such events include heating the composition to a certain threshold temperature (e.g., greater than 50 °C, preferably greater than 80°C,

more preferably greater than 100 °C), exposure to electron beam radiation, exposure to ultraviolet light, exposure to microwave energy, exposure to infrared energy, or the like. Complementary, chemically reactive moieties preferably comprise epoxy moieties and an amine moiety, as these moieties react relatively slowly with each other at room

5 temperature but very quickly when heated above a threshold temperature. This provides such compositions with reasonable shelf life and controllable crosslinking characteristics.

The crosslinkable functionality may be incorporated into S and/or D material of one or more amphipathic copolymer(s) included in the compositions of the present invention. Preferably, crosslinkable functionality is incorporated into at least the D  
10 material of at least one amphipathic copolymer included in the composition.

Advantageously, this allows the D material to be formulated with relatively low Tg constituents. D material with low Tg characteristics are desirable in some embodiments, as such material tends to have good drying characteristics, can be formulated at higher solids, are tacky and self-fixing for excellent image forming resolution, and can be fused  
15 at lower temperatures than higher Tg counterparts. Yet, such material becomes very durable, temperature resistant, and blocking resistant when crosslinked. Preferred embodiments of S material tend to have low Tg characteristics in any event, but the ability to cure the S material is also advantageous for forming more durable, temperature resistant, and blocking resistant images.

20 Such functionality is easily incorporated into the S and/or D material, as the case may be, through the use of copolymerizable monomers, oligomers, and/or polymers that contain the desired crosslinkable functionality(ies) in addition to the desired copolymerizable functionality. For example, epoxy functional, copolymerizable monomers readily incorporated into free radically polymerized S or D material include  
25 glycidyl (meth)acrylate, epoxy-9-diene, epoxy-7-octene, epoxy-6-hexene, combinations of these, and the like.

Pendant hydroxyl groups of the copolymer not only facilitate crosslinking, but also may be used to promote dispersion and interaction with the pigments in the formulation. The hydroxyl groups can be primary, secondary, or tertiary, although  
30 primary and secondary hydroxyl groups are preferred. Hydroxyl functional, copolymerizable monomers readily incorporated into free radically polymerized S or D

material include an ester of an  $\alpha$ ,  $\beta$ -unsaturated carboxylic acid with a diol, e.g., 2-hydroxyethyl (meth)acrylate, or 2-hydroxypropyl (meth)acrylate; 1,3-dihydroxypropyl-2-(meth)acrylate; 2,3-dihydroxypropyl-1-(meth)acrylate; an adduct of an  $\alpha$ ,  $\beta$ -unsaturated carboxylic acid with caprolactone; an alkanol vinyl ether such as 2-hydroxyethyl vinyl ether; 4-vinylbenzyl alcohol; allyl alcohol; p-methylol styrene; combinations of these, and/or the like.

Amine functional, copolymerizable monomers readily incorporated into free radically polymerized S or D material include DMAEMA (2-dimethylaminoethyl methacrylate), DAAM (diacetone acrylamide), combinations of these, and the like.

Isocyanate functional, copolymerizable monomers readily incorporated into free radically polymerized S or D material include TMI (dimethyl-m-isoprenyl benzyl isocyanate; ortho and para forms also may be used) IEM (isocyanatoethyl methacrylate), combinations of these, and the like.

Carboxylic acid or salt functional, copolymerizable monomers readily incorporated into free radically polymerized S or D material include methylene succinic acid, MAA (methacrylic acid), acrylic acid, 2-carboxyethyl, combinations of these, and the like.

The amount of crosslinkable functionality incorporated into the S and/or D material of an amphipathic copolymer can vary over a wide range. However, if too much is used in the S material, the desired degree of solubility of the S material could be adversely affected. If too much is used in the D material, the resultant particles may have too much of a tendency to coagulate. Balancing concerns such as these, it is preferred that each of the S and/or D material, as the case may be, incorporates 0.5 to 10, preferably about 3 to about 6 weight percent of monomers, oligomers, and/or polymers, as the case may be, bearing the desired crosslinkable functionality.

Representative embodiments of crosslinkable amphipathic copolymers of the present invention are schematically shown in Figs. 1a through 1d. In Fig. 1a, amphipathic copolymer 10 includes S material  $S_1$  and D material  $D_1$ . A crosslinkable moiety  $R_1$  is pendant from the  $S_1$  material. In Fig. 1b, amphipathic copolymer 20 includes S material  $S_1$  and D material  $D_1$ . A crosslinkable moiety  $R_1$  is pendant from the  $D_1$  material. In Fig. 1c, amphipathic copolymer 30 includes S material  $S_1$  and D material  $D_1$ .



A crosslinkable moiety  $R_1$  is pendant from both the  $S_1$  and  $D_1$  material. In Fig. 1d, amphipathic copolymer 40 includes S material  $S_1$  and D material  $D_1$ . A first crosslinkable moiety  $R_1$  is pendant from the  $S_1$  material and a second crosslinkable moiety  $R_2$  is pendant from the  $D_1$  material.

5 Preferred embodiments of the invention may comprise combinations of two or more different crosslinkable amphipathic copolymers in order to more easily achieve desired performance objective(s). For example, Fig. 2a schematically shows an organosol 50 in container 51 comprising first amphipathic copolymer 52 and second amphipathic copolymer 54 dispersed in a solvent 56. First amphipathic copolymer 52  
10 contains solvated material  $S_1$  and dispersed material  $D_1$ . A first crosslinkable moiety  $R_1$  is pendant from the  $D_1$  material. Second amphipathic copolymer 54 contains solvated material  $S_2$  and dispersed material  $D_2$ . A first crosslinkable moiety  $R_2$  is pendant from the  $D_1$  material. The  $R_1$  and  $R_2$  moieties are complementary in that these will chemically crosslink together, optionally with the assistance of one or more initiator(s), catalyst(s),  
15 crosslinking agent(s), or the like. Because each of  $R_1$  and  $R_2$  are pendant from dispersed material  $D_1$  and  $D_2$ , respectively, the complementary reactive moieties are essentially isolated from each other and will react relatively slowly with each other, if at all. However, if drying occurs above the  $T_g$  of the  $D_1$  and  $D_2$  materials, these will tend to coalesce into a film and thereby allow the  $R_1$  and  $R_2$  groups to crosslink. Pressure may  
20 also be used to bring the  $R_1$  and  $R_2$  moieties into sufficiently close proximity so as to allow crosslinking to occur. A combination of pressure and heat could also be used. The organosol 50 also is advantageously used when  $R_1$  and  $R_2$  are mutually reactive even under ambient conditions inasmuch as the  $R_1$  and  $R_2$  moieties are isolated until heated, subjected to pressure, or otherwise caused or allowed to interact. Of course, organosol 50  
25 may include other ingredients in addition to the amphipathic copolymers 52 and 54 and solvent 56, as described herein, but these are omitted for purposes of more clearly illustrating the complementary natures of the copolymer combination used in organosol 50.

Fig. 2b schematically shows another embodiment of an organosol 60 in container  
30 61 comprising first and second amphipathic copolymers 62 and 64 in a solvent 66. First amphipathic copolymer 62 contains solvated material  $S_1$  and dispersed material  $D_1$ . First

crosslinkable moieties  $R_1$  are pendant from the  $D_1$  and  $S_1$  material. Second amphipathic copolymer 64 contains solvated material  $S_2$  and dispersed material  $D_2$ . Second crosslinkable moieties  $R_2$  are pendant from the  $D_1$  and  $S_2$  material. The  $R_1$  and  $R_2$  moieties are complementary in that these will chemically crosslink together, optionally with the assistance of one or more initiator(s), catalyst(s), crosslinking agent(s), or the like. The formulation strategy of Fig. 2b is advantageously used when the  $R_1$  and  $R_2$  moieties react very slowly, and more preferably are substantially nonreactive, at room temperature or other conditions in which the liquid toner composition is likely to be stored before use to form an image, but then readily crosslink when subjected to thermal, irradiation, and/or or other curing energy. For example, when  $R_1$  comprises an epoxy moiety and  $R_2$  comprises an amine moiety, first and second amphipathic copolymers 62 and 64 are substantially non-reactive when stored, providing organosol 60 with good shelf life characteristics. But, when the organosol 60 is heated to a temperature above about 100°C, the epoxy and amine will rapidly crosslink. This is a very suitable embodiment for using low  $T_g$  S and D material to form images that are then readily cured after image forming for durability and temperature resistance.

Crosslinking of the reactive functionality pendant from amphipathic copolymers of the present invention may be achieved, either in substantial whole or in part, at any desired point(s) during the course of formulating the liquid toners, storing the toners, using the toners to form images, or the like. Preferably, crosslinking occurs subsequent to development of an image and may occur, for example, on the transfer belt, other intermediate substrate, the final substrate, or the like.

In addition to using monomers, oligomers, and/or polymer constituents that provide the amphipathic copolymers with crosslinkable functionality, other kinds of constituents may also be used to provide desired performance characteristics. For example, in order to promote hardness and abrasion resistance, a formulator may incorporate one or more free radically polymerizable monomer(s) (hereinafter "high  $T_g$  component") whose presence causes the polymerized material, or a portion thereof, to have a higher glass transition temperature,  $T_g$ , as compared to an otherwise identical material lacking such high  $T_g$  component. Preferred monomeric constituents of the high  $T_g$  component generally include monomers whose homopolymers have a  $T_g$  of at least

about 50°C, preferably at least about 60°C, and more preferably at least about 75°C in the cured state.

An exemplary class of radiation curable monomers that tend to have relatively high  $T_g$  characteristics suitable for incorporation into the high  $T_g$  component generally comprise at least one radiation curable (meth)acrylate moiety and at least one nonaromatic, alicyclic and/or nonaromatic heterocyclic moiety. Isobornyl (meth)acrylate is a specific example of one such monomer. A cured, homopolymer film formed from isobornyl acrylate, for instance, has a  $T_g$  of 110°C. The monomer itself has a molecular weight of 222 g/mole, exists as a clear liquid at room temperature, has a viscosity of 9 centipoise at 25°C, and has a surface tension of 31.7 dynes/cm at 25°C. Additionally, 1,6-Hexanediol di(meth)acrylate is another example of a monomer with high  $T_g$  characteristics.

Trimethyl cyclohexyl methacrylate (TCHMA) is another example of a high  $T_g$  monomer useful in the practice of the present invention. TCHMA has a  $T_g$  of 125°C and tends to be soluble in oleophilic solvents. Consequently, TCHMA is easily incorporated into S material. However, if used in limited amounts so as not to unduly impair the insolubility characteristics of D material, some TCHMA may also be incorporated into D the material.

The advantages of incorporating High  $T_g$  Monomers into the copolymer are further described in assignee's co-pending U.S. Patent Application titled **ORGANOSOL INCLUDING HIGH  $T_g$  AMPHIPATHIC COPOLYMERIC BINDER AND LIQUID TONERS FOR ELECTROPHOTOGRAPHIC APPLICATIONS**, bearing Attorney Docket No. SAM0005/US, filed November 12, 2002, in the names of **James A. Baker et al.** The advantages of incorporating Soluble High  $T_g$  Monomer into the copolymer are further described in assignee's co-pending U.S. Patent Application titled **ORGANOSOL INCLUDING AMPHIPATHIC COPOLYMERIC BINDER MADE WITH SOLUBLE HIGH  $T_g$  MONOMER AND LIQUID TONERS FOR ELECTROPHOTOGRAPHIC APPLICATIONS**, bearing Attorney Docket No. SAM0006/US, filed November 12, 2002, in the names of **James A. Baker et al.** Both of these co-pending patent applications are hereby incorporated herein by reference in their entirety. Nitrile functionality may be advantageously incorporated into the copolymer for

a variety of reasons, including improved durability, enhanced compatibility with visual enhancement additive(s), e.g., colorant particles, and the like. In order to provide a copolymer having pendant nitrile groups, one or more nitrile functional monomers can be used. Representative examples of such monomers include (meth)acrylonitrile,  $\beta$ -  
5 cyanoethyl-(meth)acrylate, 2-cyanoethoxyethyl (meth)acrylate, p-cyanostyrene, p-(cyanomethyl)styrene, N-vinylpyrrolidinone, and the like.

In certain preferred embodiments, polymerizable crystallizable compounds, e.g. crystalline monomer(s) are chemically incorporated into the copolymer. Above the melting point of the crystalline material, the crystalline material helps to suppress the T<sub>g</sub>  
10 of an amphipathic copolymer, allowing lower fusing temperatures to be used for image forming. Yet, below the melting point of the crystalline materials, the high T<sub>g</sub> characteristics of the amphipathic copolymer are essentially unaffected as a practical matter. The term "crystalline monomer" refers to a monomer whose homopolymeric analog is capable of independently and reversibly crystallizing at or above room  
15 temperature (e.g., 22°C).

If used, one or more of these crystalline monomers may be incorporated into the D material of the copolymer. Suitable crystalline monomers include alkyl(meth)acrylates where the alkyl chain contains more than 13 carbon atoms (e.g. tetradecyl(meth)acrylate, pentadecyl(meth)acrylate, hexadecyl(meth)acrylate, heptadecyl(meth)acrylate,  
20 octadecyl(meth)acrylate, etc). Other suitable crystalline monomers whose homopolymers have melting points above 22°C include aryl acrylates and methacrylates; high molecular weight alpha olefins; linear or branched long chain alkyl vinyl ethers or vinyl esters; long chain alkyl isocyanates; unsaturated long chain polyesters, polysiloxanes and polysilanes; polymerizable natural waxes with melting points above 22°C, polymerizable synthetic  
25 waxes with melting points above 22°C, and other similar type materials known to those skilled in the art.

It will be understood by those skilled in the art that blocking resistance can be observed at temperatures above room temperature but below the crystallization temperature of the polymer portion incorporating the crystalline monomers or other  
30 polymerizable crystallizable compound. Many crystalline monomers tend to be soluble in oleophilic solvents commonly used as liquid carrier material(s) in an organosol. Thus,

crystalline material is relatively easily incorporated into S material without impacting desired solubility characteristics. However, if too much of such crystalline material were to be incorporated into D material, the resultant D material may tend to be too soluble in the organosol. Yet, so long as the amount of soluble, crystalline material in the D material is limited, some amount of crystalline material may be advantageously incorporated into the D material without unduly impacting the desired insolubility characteristics. Thus, when present in the D material, the crystalline material is preferably provided in an amount of up to about 30%, more preferably up to about 20%, most preferably up to about 5% to 10% of the total D material incorporated into the copolymer.

When crystalline monomers or PCC's are chemically incorporated into the D material, suitable co-polymerizable compounds to be used in combination with the PCC include monomers (including other PCC's) such as 2-ethylhexyl acrylate, 2-ethylhexyl (methacrylate), lauryl acrylate, lauryl methacrylate, octadecyl acrylate, octadecyl(methacrylate), isobornyl acrylate, isobornyl (methacrylate), hydroxy(ethylmethacrylate), and other acrylates and methacrylates.

The use of crystalline materials in amphipathic copolymers to form liquid and dry toner compositions is further described in co-pending U.S. Patent Application titled **ORGANOSOL LIQUID TONER INCLUDING AMPHIPATHIC COPOLYMERIC BINDER HAVING CRYSTALLINE COMPONENT**, bearing Attorney Docket No. **SAM0004/P1**, and filed on November 12, 2002, in the names of **James A. Baker et al.**

Multifunctional free radically reactive materials may also be incorporated into amphipathic copolymers of the present invention, if desired, to enhance one or more properties of the resultant toner particles, including crosslink density, hardness, tackiness, mar resistance, or the like. Examples of such higher functional, monomers include ethylene glycol di(meth)acrylate, hexanediol di(meth)acrylate, triethylene glycol di(meth)acrylate, tetraethylene glycol di(meth)acrylate, trimethylolpropane tri(meth)acrylate, ethoxylated trimethylolpropane tri(meth)acrylate, glycerol tri(meth)acrylate, pentaerythritol tri(meth)acrylate, pentaerythritol tetra(meth)acrylate, and neopentyl glycol di(meth)acrylate, divinyl benzene, combinations of these, and the like.

Suitable free radically reactive oligomer and/or polymeric materials for use in the present invention include, but are not limited to, (meth)acrylated urethanes (i.e., urethane (meth)acrylates), (meth)acrylated epoxies (i.e., epoxy (meth)acrylates), (meth)acrylated polyesters (i.e., polyester (meth)acrylates), (meth)acrylated (meth)acrylics,  
5 (meth)acrylated silicones, (meth)acrylated polyethers (i.e., polyether (meth)acrylates), vinyl (meth)acrylates, and (meth)acrylated oils.

Copolymers of the present invention can be prepared by free-radical polymerization methods known in the art, including but not limited to bulk, solution, and dispersion polymerization methods. The resultant copolymers may have a variety of  
10 structures including linear, branched, three dimensionally networked, graft-structured, combinations thereof, and the like. A preferred embodiment is a graft copolymer comprising one or more oligomeric and/or polymeric arms attached to an oligomeric or polymeric backbone. In graft copolymer embodiments, the S portion or D portion materials, as the case may be, may be incorporated into the arms and/or the backbone.

Any number of reactions known to those skilled in the art may be used to prepare  
15 a free radically polymerized copolymer having a graft structure. Common grafting methods include random grafting of polyfunctional free radicals; copolymerization of monomers with macromonomers; ring-opening polymerizations of cyclic ethers, esters, amides or acetals; epoxidations; reactions of hydroxyl or amino chain transfer agents with  
20 terminally-unsaturated end groups; esterification reactions (i.e., glycidyl methacrylate undergoes tertiary-amine catalyzed esterification with methacrylic acid); and condensation polymerization.

Representative methods of forming graft copolymers are described in U.S. Pat. Nos. 6,255,363; 6,136,490; and 5,384,226; and Japanese Published Patent Document No.  
25 05-119529, incorporated herein by reference. Representative examples of grafting methods are also described in sections 3.7 and 3.8 of Dispersion Polymerization in Organic Media, K.E.J. Barrett, ed., (John Wiley; New York, 1975) pp. 79-106, also incorporated herein by reference.

Representative examples of grafting methods also may use an anchoring group to  
30 facilitate anchoring. The function of the anchoring group is to provide a covalently bonded link between the core part of the copolymer (the D material) and the soluble shell

component (the S material). Suitable monomers containing anchoring groups include: adducts of alkenylazlactone comonomers with an unsaturated nucleophile containing hydroxy, amino, or mercaptan groups, such as 2-hydroxyethylmethacrylate, 3-hydroxypropylmethacrylate, 2-hydroxyethylacrylate, pentaerythritol triacrylate, 4-hydroxybutylvinylether, 9-octadecen-1-ol, cinnamyl alcohol, allyl mercaptan, methallylamine; and azlactones, such as 2-alkenyl-4,4-dialkylazlactone.

The preferred methodology described above accomplishes grafting via attaching an ethylenically-unsaturated isocyanate (e.g., dimethyl-m-isopropenyl benzylisocyanate, TMI, available from CYTEC Industries, West Paterson, NJ; or isocyanatoethyl methacrylate, also known as IEM) to hydroxyl groups in order to provide free radically reactive anchoring groups.

A preferred method of forming a graft copolymer of the present invention involves three reaction steps that are carried out in a suitable substantially nonaqueous liquid carrier in which resultant S material is soluble while D material is dispersed or insoluble.

In a first preferred step, a hydroxyl functional, free radically polymerized oligomer or polymer is formed from one or more monomers, wherein at least one of the monomers has pendant hydroxyl functionality. Preferably, the hydroxyl functional monomer constitutes about 1 to about 30, preferably about 2 to about 10 percent, most preferably 3 to about 5 percent by weight of the monomers used to form the oligomer or polymer of this first step. This first step is preferably carried out via solution polymerization in a substantially nonaqueous solvent in which the monomers and the resultant polymer are soluble. For instance, using the Hildebrand solubility data in Table 1, monomers such as octadecyl methacrylate, octadecyl acrylate, lauryl acrylate, and lauryl methacrylate are suitable for this first reaction step when using an oleophilic solvent such as heptane or the like.

In a second reaction step, all or a portion of the hydroxyl groups of the soluble polymer are catalytically reacted with an ethylenically unsaturated aliphatic isocyanate (e.g. meta-isopropenyldimethylbenzyl isocyanate commonly known as TMI or isocyanatoethyl methacrylate, commonly known as IEM) to form pendant free radically polymerizable functionality which is attached to the oligomer or polymer via a

polyurethane linkage. This reaction can be carried out in the same solvent, and hence the same reaction vessel, as the first step. The resultant double-bond functionalized polymer generally remains soluble in the reaction solvent and constitutes the S portion material of the resultant copolymer, which ultimately will constitute at least a portion of the

5 solvatable portion of the resultant triboelectrically charged particles.

The resultant free radically reactive functionality provides grafting sites for attaching D material and optionally additional S material to the polymer. In a third step, these grafting site(s) are used to covalently graft such material to the polymer via reaction with one or more free radically reactive monomers, oligomers, and or polymers that are initially soluble in the solvent, but then become insoluble as the molecular weight of the graft copolymer. For instance, using the Hildebrand solubility parameters in Table 1, monomers such as e.g. methyl (meth)acrylate, ethyl (meth)acrylate, t-butyl methacrylate and styrene are suitable for this third reaction step when using an oleophilic solvent such as heptane or the like.

15 The product of the third reaction step is generally an organosol comprising the resultant copolymer dispersed in the reaction solvent, which constitutes a substantially nonaqueous liquid carrier for the organosol. At this stage, it is believed that the copolymer tends to exist in the liquid carrier as discrete, monodisperse particles having dispersed (e.g., substantially insoluble, phase separated) portion(s) and solvated (e.g.,  
20 substantially soluble) portion(s). As such, the solvated portion(s) help to sterically stabilize the dispersion of the particles in the liquid carrier. It can be appreciated that the copolymer is thus advantageously formed in the liquid carrier in situ.

Before further processing, the copolymer particles may remain in the reaction solvent. Alternatively, the particles may be transferred in any suitable way into fresh  
25 solvent that is the same or different so long as the copolymer has solvated and dispersed phases in the fresh solvent. In either case, the resulting organosol is then converted into toner particles by mixing the organosol with at least one visual enhancement additive. Optionally, one or more other desired ingredients also can be mixed into the organosol before and/or after combination with the visual enhancement particles. During such  
30 combination, it is believed that ingredients comprising the visual enhancement additive and the copolymer will tend to self-assemble into composite particles having a structure



wherein the dispersed phase portions generally tend to associate with the visual enhancement additive particles (for example, by physically and/or chemically interacting with the surface of the particles), while the solvated phase portions help promote dispersion in the carrier.

5           If more than one kind of amphipathic copolymer is incorporated into a liquid toner, these can be made separately and then mixed together as well. This can be done prior to packaging, for instance, if the crosslinkable functionality on one or more of the amphipathic copolymers is sufficiently nonreactive under the expected storage conditions. Alternatively, if the crosslinkable functionality of two or more amphipathic  
10 copolymers are too reactive under expected storage conditions, the components may be separately packaged and then combined or serially dispensed at the time of use.

The optional visual enhancement additive(s) generally may include any one or more fluid and/or particulate materials that provide a desired visual effect when toner particles incorporating such materials are printed onto a receptor. Examples include one  
15 or more colorants, fluorescent materials, pearlescent materials, iridescent materials, metallic materials, flip-flop pigments, silica, polymeric beads, reflective and non-reflective glass beads, mica, combinations of these, and the like. The amount of visual enhancement additive incorporated into the toner particles may vary over a wide range. In representative embodiments, a suitable weight ratio of copolymer to visual  
20 enhancement additive is from 1/1 to 20/1, preferably from 2/1 to 10/1 and most preferably from 4/1 to 8/1.

Useful colorants are well known in the art and include materials listed in the Colour Index, as published by the Society of Dyers and Colourists (Bradford, England), including dyes, stains, and pigments. Preferred colorants are pigments which may be  
25 combined with ingredients comprising the copolymer to interact with the D portion of the copolymer to form dry toner particles with structure as described herein, are at least nominally insoluble in and nonreactive with the carrier liquid, and are useful and effective in making visible the latent electrostatic image. It is understood that the visual enhancement additive(s) may also interact with each other physically and/or chemically,  
30 forming aggregations and/or agglomerates of visual enhancement additives that also interact with the D portion of the copolymer. Examples of suitable colorants include:

phthalocyanine blue (C.I. Pigment Blue 15:1, 15:2, 15:3 and 15:4), monoarylide yellow (C.I. Pigment Yellow 1, 3, 65, 73 and 74), diarylide yellow (C.I. Pigment Yellow 12, 13, 14, 17 and 83), arylamide (Hansa) yellow (C.I. Pigment Yellow 10, 97, 105 and 111), isoindoline yellow (C.I. Pigment Yellow 138), azo red (C.I. Pigment Red 3, 17, 22, 23, 38, 48:1, 48:2, 52:1, and 52:179), quinacridone magenta (C.I. Pigment Red 122, 202 and 209), laked rhodamine magenta (C.I. Pigment Red 81:1, 81:2, 81:3, and 81:4), and black pigments such as finely divided carbon (Cabot Monarch 120, Cabot Regal 300R, Cabot Regal 350R, Vulcan X72, and Aztech ED 8200), and the like.

In addition to the visual enhancement additive, other additives optionally can be formulated into the liquid toner composition. A particularly preferred additive comprises at least one charge control agent (CCA, charge control additive or charge director). The charge control agent, also known as a charge director, can be included as a separate ingredient and/or included as one or more functional moiety(ies) of the S and/or D material incorporated into the amphipathic copolymer. The charge control agent acts to enhance the chargeability and/or impart a charge to the toner particles. Toner particles can obtain either positive or negative charge depending upon the combination of particle material and charge control agent.

The charge control agent can be incorporated into the toner particles using a variety of methods, such as copolymerizing a suitable monomer with the other monomers used to form the copolymer, chemically reacting the charge control agent with the toner particle, chemically or physically adsorbing the charge control agent onto the toner particle (resin or pigment), or chelating the charge control agent to a functional group incorporated into the toner particle. One preferred method is via a functional group built into the S material of the copolymer.

The charge control agent acts to impart an electrical charge of selected polarity onto the toner particles. Any number of charge control agents described in the art can be used. For example, the charge control agent can be provided in the form of metal salts consisting of polyvalent metal ions and organic anions as the counterion. Suitable metal ions include, but are not limited to, Ba(II), Ca(II), Mn(II), Zn(II), Zr(IV), Cu(II), Al(III), Cr(III), Fe(II), Fe(III), Sb(III), Bi(III), Co(II), La(III), Pb(II), Mg(II), Mo(III), Ni(II), Ag(I), Sr(II), Sn(IV), V(V), Y(III), and Ti(IV). Suitable organic anions include

carboxylates or sulfonates derived from aliphatic or aromatic carboxylic or sulfonic acids, preferably aliphatic fatty acids such as stearic acid, behenic acid, neodecanoic acid, diisopropylsalicylic acid, octanoic acid, abietic acid, naphthenic acid, lauric acid, tallic acid, and the like.

5 Preferred negative charge control agents are lecithin and basic barium petronate. Preferred positive charge control agents include metallic carboxylates (soaps), for example, as described in U.S. Pat. No. 3,411,936 (incorporated herein by reference). A particularly preferred positive charge control agent is zirconium tetraoctoate (available as Zirconium HEX-CEM from OMG Chemical Company, Cleveland, OH).

10 The preferred charge control agent levels for a given toner formulation will depend upon a number of factors, including the composition of the S portion and the organosol, the molecular weight of the organosol, the particle size of the organosol, the D:S ratio of the polymeric binder, the pigment used in making the toner composition, and the ratio of organosol to pigment. In addition, preferred charge control agent levels will  
15 depend upon the nature of the electrophotographic imaging process. The level of charge control agent can be adjusted based upon the parameters listed herein, as known in the art. The amount of the charge control agent, based on 100 parts by weight of the toner solids, is generally in the range of 0.01 to 10 parts by weight, preferably 0.1 to 5 parts by weight.

20 The conductivity of a liquid toner composition can be used to describe the effectiveness of the toner in developing electrophotographic images. A range of values from  $1 \times 10^{-11}$  mho/cm to  $3 \times 10^{-10}$  mho/cm is considered advantageous to those of skill in the art. High conductivities generally indicate inefficient association of the charges on the toner particles and is seen in the low relationship between current density and toner  
25 deposited during development. Low conductivities indicate little or no charging of the toner particles and lead to very low development rates. The use of charge control agents matched to adsorption sites on the toner particles is a common practice to ensure sufficient charge associates with each toner particle.

Other additives may also be added to the formulation in accordance with  
30 conventional practices. These include one or more of UV stabilizers, mold inhibitors,

bactericides, fungicides, antistatic agents, gloss modifying agents, other polymer or oligomer material, antioxidants, and the like.

5 The particle size of the resultant charged toner particles can impact the imaging, fusing, resolution, and transfer characteristics of the toner composition incorporating such particles. Preferably, the volume mean particle diameter (determined with laser diffraction) of the particles is in the range of about 0.05 to about 50.0 microns, more preferably in the range of about 3 to about 10 microns, most preferably in the range of about 1.5 to about 5 microns.

10 In electrophotographic and electrographic processes, an electrostatic image is formed on the surface of a photoreceptive element or dielectric element, respectively. The photoreceptive element or dielectric element may be an intermediate transfer drum or belt or the substrate for the final toned image itself, as described by Schmidt, S. P. and Larson, J. R. in Handbook of Imaging Materials Diamond, A. S., Ed: Marcel Dekker: New York; Chapter 6, pp 227-252, and U.S. Pat. Nos. 4,728,983, 4,321,404, and  
15 4,268,598.

In electrography, a latent image is typically formed by (1) placing a charge image onto the dielectric element (typically the receiving substrate) in selected areas of the element with an electrostatic writing stylus or its equivalent to form a charge image, (2) applying toner to the charge image, and (3) fixing the toned image. An example of this  
20 type of process is described in U.S. Pat. No. 5,262,259. Images formed by the present invention may be of a single color or a plurality of colors. Multicolor images can be prepared by repetition of the charging and toner application steps.

In electrophotography, the electrostatic image is typically formed on a drum or belt coated with a photoreceptive element by (1) uniformly charging the photoreceptive  
25 element with an applied voltage, (2) exposing and discharging portions of the photoreceptive element with a radiation source to form a latent image, (3) applying a toner to the latent image to form a toned image, and (4) transferring the toned image through one or more steps to a final receptor sheet. In some applications, it is sometimes desirable to fix the toned image using a heated pressure roller or other fixing methods  
30 known in the art.

While the electrostatic charge of either the toner particles or photoreceptive element may be either positive or negative, electrophotography as employed in the present invention is preferably carried out by dissipating charge on a positively charged photoreceptive element. A positively-charged toner is then applied to the regions in  
5 which the positive charge was dissipated using a dry toner development technique.

The substrate for receiving the image from the photoreceptive element can be any commonly used receptor material, such as paper, coated paper, polymeric films and primed or coated polymeric films. Polymeric films include polyesters and coated polyesters, polyolefins such as polyethylene or polypropylene, plasticized and  
10 compounded polyvinyl chloride (PVC), acrylics, polyurethanes, polyethylene/acrylic acid copolymer, and polyvinyl butyrals. The polymer film may be coated or primed, e.g. to promote toner adhesion.

Fig. 3 schematically illustrates how the principles of the present invention may be incorporated into structures with tamper-resistant images. For purposes of illustration,  
15 Fig. 3 shows a cross-section of an identification device 70 such as an employee badge, drivers license, or the like. Device 70 includes an image 72 formed on a substrate 74. Image 72 incorporates a liquid toner composition of the present invention having one or more types of crosslinkable functionality denoted by the designation R1. Substrate 74 includes complementary crosslinkable functionality denoted by the designation R2.  
20 Transparent coverlay 76 overlies image 72 and includes crosslinkable functionality R3 that is complementary to R1 and/or R2, preferably at least R1. The crosslinkable functionality R2 may be the same or different than one or both of R1 and R2, depending upon the nature of the crosslinkable functionality. In a particularly preferred embodiment, R2 and R3 are the same and are complementary to R1. When the device 70  
25 is subjected to crosslinking, the image 72 becomes covalently linked to substrate 74 and/or coverlay 76, as the case may be. When device 70 is pulled apart, the image 72 will be split or otherwise destroyed, making it difficult for the image 72 of device 70 to be modified after the manufacture of device 70. Authenticity and tamper-resistance of image 72 are thus enhanced.

30 These and other aspects of the present invention are demonstrated in the illustrative examples that follow.

In the practice of the present invention, as shown in the following examples, percent solids of the copolymer solutions and the organosol and ink dispersions were determined gravimetrically using the Halogen Lamp Drying Method using a halogen lamp drying oven attachment to a precision analytical balance (Mettler Instruments, Inc.,  
5 Highstown, N.J.). Approximately two grams of sample were used in each determination of percent solids using this sample dry down method.

In the practice of the invention, molecular weight is normally expressed in terms of the weight average molecular weight, while molecular weight polydispersity is given by the ratio of the weight average molecular weight to the number average molecular  
10 weight. Molecular weight parameters were determined with gel permeation chromatography (GPC) using tetrahydrofuran as the carrier solvent. Absolute weight average molecular weight were determined using a Dawn DSP-F light scattering detector (Wyatt Technology Corp., Santa Barbara, Calif.), while polydispersity was evaluated by ratioing the measured weight average molecular weight to a value of number average  
15 molecular weight determined with an Optilab 903 differential refractometer detector (Wyatt Technology Corp., Santa Barbara, Calif.).

Organosol and toner particle size distributions were determined by the Laser Diffraction Light Scattering Method using a Horiba LA-900 laser diffraction particle size analyzer (Horiba Instruments, Inc., Irvine, Calif.). Samples are diluted approximately  
20 1/500 by volume and sonicated for one minute at 150 watts and 20 kHz prior to measurement. Particle size was expressed as both a number mean diameter ( $D_n$ ) and a volume mean diameter ( $D_v$ ) and in order to provide an indication of both the fundamental (primary) particle size and the presence of aggregates or agglomerates.

The liquid toner conductivity (bulk conductivity,  $k_b$ ) was determined at  
25 approximately 18 Hz using a Scientifica Model 627 conductivity meter (Scientifica Instruments, Inc., Princeton, N.J.). In addition, the free (liquid dispersant) phase conductivity ( $k_f$ ) in the absence of toner particles was also determined. Toner particles were removed from the liquid medium by centrifugation at 5°C for 1-2 hours at 6,000 rpm (6,110 relative centrifugal force) in a Jouan MR1822 centrifuge (Winchester, VA).  
30 The supernatant liquid was then carefully decanted, and the conductivity of this liquid was measured using a Scientifica Model 627 conductance meter. The percentage of free

phase conductivity relative to the bulk toner conductivity was then determined as 100% ( $k_f/k_b$ ).

Toner particle electrophoretic mobility (dynamic mobility) was measured using a Matec MBS-8000 Electrokinetic Sonic Amplitude Analyzer (Matec Applied Sciences, Inc., Hopkinton, MA). Unlike electrokinetic measurements based upon microelectrophoresis, the MBS-8000 instrument has the advantage of requiring no dilution of the toner sample in order to obtain the mobility value. Thus, it was possible to measure toner particle dynamic mobility at solids concentrations actually preferred in printing. The MBS-8000 measures the response of charged particles to high frequency (1.2 MHz) alternating (AC) electric fields. In a high frequency AC electric field, the relative motion between charged toner particles and the surrounding dispersion medium (including counter-ions) generates an ultrasonic wave at the same frequency of the applied electric field. The amplitude of this ultrasonic wave at 1.2 MHz can be measured using a piezoelectric quartz transducer; this electrokinetic sonic amplitude (ESA) is directly proportional to the low field AC electrophoretic mobility of the particles. The particle zeta potential can then be computed by the instrument from the measured dynamic mobility and the known toner particle size, liquid dispersant viscosity, and liquid dielectric constant.

The charge per mass measurement (Q/M) was measured using an apparatus that consists of a conductive metal plate, a glass plate coated with Indium Tin Oxide (ITO), a high voltage power supply, an electrometer, and a personal computer (PC) for data acquisition. A 1% solution of ink was placed between the conductive plate and the ITO coated glass plate. An electrical potential of known polarity and magnitude was applied between the ITO coated glass plate and the metal plate, generating a current flow between the plates and through wires connected to the high voltage power supply. The electrical current was measured 100 times a second for 20 seconds and recorded using the PC. The applied potential causes the charged toner particles to migrate towards the plate (electrode) having opposite polarity to that of the charged toner particles. By controlling the polarity of the voltage applied to the ITO coated glass plate, the toner particles may be made to migrate to that plate.

The ITO coated glass plate was removed from the apparatus and placed in an oven for approximately 30 minutes at 50°C to dry the plated ink completely. After drying, the ITO coated glass plate containing the dried ink film was weighed. The ink was then removed from the ITO coated glass plate using a cloth wipe impregnated with Norpar™ 12, and the clean ITO glass plate was weighed again. The difference in mass between the dry ink coated glass plate and the clean glass plate is taken as the mass of ink particles (m) deposited during the 20 second plating time. The electrical current values were used to obtain the total charge carried by the toner particles (Q) over the 20 seconds of plating time by integrating the area under a plot of current vs. time using a curve-fitting program (e.g. TableCurve 2D from Systat Software Inc.). The charge per mass (Q/m) was then determined by dividing the total charge carried by the toner particles by the dry plated ink mass .

In the following examples, toner was printed onto final image receptors using the following methodology (referred to in the Examples as the Liquid Electrophotographic Printing Method):

A light sensitive temporary image receptor (organic photoreceptor or “OPC”) was charged with a uniform positive charge of approximately 850 volts. The positively charged surface of the OPC was image-wise irradiated with a scanning infrared laser module in order to reduce the charge wherever the laser struck the surface. Typical charge-reduced values were between 50 volts and 100 volts.

A developer apparatus was then utilized to apply the toner particles to the OPC surface. The developer apparatus included the following elements: a conductive rubber developer roll in contact with the OPC, liquid toner, a conductive deposition roll, an insulative foam cleaning roll in contact with developer roll surface, and a conductive skiving blade (skive) in contact with the developer roll. The contact area between the developer roll and the OPC is referred to as the “developing nip.” The developer roll and conductive deposition roll were both partially suspended in the liquid toner. The developer roll delivered liquid toner to the OPC surface, while the conductive deposition roll was positioned with its roll axis parallel to the developer roll axis and its surface arranged to be approximately 150 microns from the surface of the developer roll, thereby forming a deposition gap.



During development, toner was initially transferred to the developer roll surface by applying a voltage of approximately 500 volts to the conductive developer roll and applying a voltage of 600 volts to the deposition roll. This created a 100-volt potential between the developer roll and the deposition roll so that in the deposition gap, toner particles (which were positively charged) migrated to the surface of the developer roll and remained there as the developer roll surface exited from the liquid toner into the air.

The conductive metal skive was biased to at least 600 volts (or more) and skived liquid toner from the surface of the developer roll without scraping off the toner layer that was deposited in the deposition gap. The developer roll surface at this stage contained a uniformly thick layer of toner at approximately 25% solids. As this toner layer passed through the developing nip, toner was transferred from the developer roll surface to the OPC surface in all the discharged areas of the OPC (the charge image), since the toner particles were positively charged. At the exit of the developing nip, the OPC contained a toner image and the developer roll contained a negative of that toner image which was subsequently cleaned from the developer roll surface by encountering the rotating foam cleaning roll.

The developed latent image (toned image) on the photoreceptor was subsequently transferred to the final image receptor without film formation of the toner on the OPC. Transfer was effected either directly to the final image receptor, or indirectly using an electrostatically-assisted offset transfer to an Intermediate Transfer Belt (ITB), with subsequent electrostatically-assisted offset transfer to the final image receptor. Smooth, clay coated papers were preferred final image receptors for direct transfer of a non-film formed toner from the photoreceptor, while plain, uncoated 20 pound bond paper was a preferred final image receptor for offset transfer using an electrostatic assist. Electrostatically-assisted transfer of non film-formed toner was most effective when the transfer potential (potential difference between the toner on the OPC and the paper back-up roller for direct transfer; or potential difference between the toner on the OPC and the ITB for offset transfer) was maintained in the range of 200-1000 V or 800-2000 V, respectively.

## **Materials**

The following abbreviations are used in the examples:

BHA: Behenyl acrylate (a PCC available from Ciba Specialty Chemical Co., Suffolk, VA)

5 BMA: Butyl methacrylate (available from Aldrich Chemical Co., Milwaukee, WI)

DAAM: Diacetone acrylamide (Aldrich Chemical Co., Milwaukee, WI)

DMAEMA: 2-Dimethylaminoethyl methacrylate (Aldrich Chemical Co., Milwaukee, WI)

EMA: Ethyl methacrylate (available from Aldrich Chemical Co., Milwaukee, WI)

10 Exp 61: Amine-functional silicone wax (a PCC available from Genesee Polymer Corporation, Flint, MI)

GMA: Glycidyl methacrylate (Aldrich Chemical Co., Milwaukee, WI)

HEMA: 2-Hydroxyethyl methacrylate (available from Aldrich Chemical Co., Milwaukee, WI)

15 LMA: Lauryl methacrylate (available from Aldrich Chemical Co., Milwaukee, WI)

MAA: Methacrylate acid (Aldrich Chemical Co., Milwaukee, WI)

ODA: Octadecyl acrylate (a PCC available Aldrich Chemical Co., Milwaukee, WI)

TCHMA: Trimethyl cyclohexyl methacrylate (available from Ciba Specialty Chemical Co., Suffolk, Virginia)

20 St: Styrene (available from Aldrich Chemical Co., Milwaukee, WI)

TMI: Dimethyl-m-isopropenyl benzyl isocyanate (available from CYTEC Industries, West Paterson, NJ)

AIBN: Azobisisobutyronitrile (an initiator available as VAZO-64 from DuPont Chemical Co., Wilmington, DE)

25 V-601: Dimethyl 2, 2'-azobisisobutyrate (an initiator available as V-601 from WAKO Chemicals U.S.A., Richmond, VA)

DBTDL: Dibutyl tin dilaurate (a catalyst available from Aldrich Chemical Co., Milwaukee, WI)

Zirconium HEX-CEM: (metal soap, zirconium tetraoctoate, available from OMG

30 Chemical Company, Cleveland, OH)

In the following examples, the compositional details of each copolymer will be summarized by ratioing the weight percentages of monomers used to create the copolymer. The grafting site composition is expressed as a weight percentage of the monomers comprising the copolymer or copolymer precursor, as the case may be. For example, a graft stabilizer (precursor to the S portion of the copolymer) designated TCHMA/HEMA-TMI (97/3-4.7) is made by copolymerizing, on a relative basis, 97 parts by weight TCHMA and 3 parts by weight HEMA, and this hydroxy functional polymer was reacted with 4.7 parts by weight of TMI.

Similarly, a graft copolymer organosol designated TCHMA/HEMA-TMI//EMA (97-3-4.7//100) is made by copolymerizing the designated graft stabilizer (TCHMA/HEMA-TMI (97/3-4.7)) (S portion or shell) with the designated core monomer EMA (D portion or core, 100% EMA) at a specified ratio of D/S (core/shell) determined by the relative weights reported in the examples.

#### **Example 1 (Comparative)**

A 5000 ml 3-neck round flask equipped with a condenser, a thermocouple connected to a digital temperature controller, a nitrogen inlet tube connected to a source of dry nitrogen and a magnetic stirrer, was charged with a mixture of 2561 g of Norpar<sup>TM</sup> 15, 849 g of LMA, 26.8 g of 98% HEMA and 8.75 g of V601. While stirring the mixture, the reaction flask was purged with dry nitrogen for 30 minutes at flow rate of approximately 2 liters/minute. A hollow glass stopper was then inserted into the open end of the condenser and the nitrogen flow rate was reduced to approximately 0.5 liters/minute. The mixture was heated to 70° C for 16 hours. The conversion was quantitative.

The mixture was then heated to 90° C and held at that temperature for 1 hour to destroy any residual V601, and then was cooled back to 70° C. The nitrogen inlet tube was then removed, and 13.6 g of 95% DBTDL were added to the mixture, followed by 41.1 g of TMI. The TMI was added drop wise over the course of approximately 5 minutes while stirring the reaction mixture. The nitrogen inlet tube was replaced, the hollow glass stopper in the condenser was removed, and the reaction flask was purged with dry nitrogen for 30 minutes at a flow rate of approximately 2 liters/minute. The

hollow glass stopper was reinserted into the open end of the condenser and the nitrogen flow rate was reduced to approximately 0.5 liters/min. The mixture was allowed to react at 70° C for 6 hours, at which time the conversion was quantitative.

The mixture was then cooled to room temperature. The cooled mixture was a viscous, transparent liquid containing no visible insoluble mater. The percent solids of the liquid mixture was determined to be 25.64% using the halogen drying method described above. Subsequent determination of molecular weight was made using the GPC method described above; the copolymer had a  $M_w$  of 231,350 Da and  $M_w/M_n$  of 3.2 based on two independent measurements. The product was a copolymer of LMA and HEMA containing random side chains of TMI and was designed herein as LMA/HEMA-TMI (97/3-4.7% w/w) and suitable for making an organosol containing non-reactive groups in the shell.

### Example 2

Using the method and apparatus of Example 1, 2561 g of Norpar<sup>TM</sup>15, 823 g of LMA, 26 g of DAAM, 26.8 g of 98% HEMA and 8.75 g of V601 were combined and resulting mixture reacted at 70° C for 16 hours. The mixture was then heated to 90° C for 1 hour to destroy any residual V601, and then was cooled back to 70° C. To the cooled mixture was then added 13.6 g of 95% DBTDL and 41.1 g of TMI. The TMI was added drop wise over the course of approximately 5 minutes while stirring the reaction mixture. Following the procedure of Example 1, the mixture was reacted at 70° C for approximately 6 hours at which time the reaction was quantitative. The mixture was then cooled to room temperature. The cooled mixture was a viscous, transparent solution, containing no visible insoluble mater.

The percent solids of the liquid mixture was determined to be 24.47% using the halogen drying method described above. Subsequent determination of molecular weight was made using the GPC method described above; the copolymer had a  $M_w$  of 278,800 Da and  $M_w/M_n$  of 2.50 based upon two independent measurements. The product was a copolymer of LMA, DAAM and HEMA containing random side chains of TMI and was designed herein as LMA/DAAM/HEMA-TMI (94/3/3-4.7% w/w) and was suitable for making an organosol containing secondary amine reactive groups in the shell.

### Example 3

Using the method and apparatus of Example 1, 2561 g of Norpar™ 15, 823 g of LMA, 26 g of MAA, 26.8 g of 98% HEMA and 8.75 g of V601 were combined and  
5 resulting mixture reacted at 70° C for 16 hours. The mixture was then heated to 90° C for 1 hour to destroy any residual V601, and then was cooled back to 70°C. To the cooled mixture was then added 13.6 g of 95% DBTDL and 41.1 g of TMI. The TMI was added drop wise over the course of approximately 5 minutes while stirring the reaction mixture. Following the procedure of Example 1, the mixture was reacted at 70° C for  
10 approximately 6 hours at which time the reaction was quantitative. The mixture was then cooled to room temperature. The cooled mixture was viscous, transparent solution, containing no visible insoluble mater.

The percent solids of the liquid mixture was determined to be 25.10% using the halogen drying method described above. Subsequent determination of molecular weight  
15 was made using the GPC method described above; the copolymer had a  $M_w$  of 330,300 Da and  $M_w/M_n$  of 2.34 based upon two independent measurements. The product was a copolymer of LMA, MAA and HEMA containing random side chains of TMI and was designed herein as LMA/MAA/HEMA-TMI (94/3/3-4.7% w/w) and was suitable for making an organosol containing carboxyl reactive groups in the shell.

### Example 4

Using the method and apparatus of Example 1, 2561 g of Norpar™ 15, 796 g of LMA, 53 g of GMA, 26.8 g of 98% HEMA and 8.75 g of V601 were combined and  
25 resulting mixture reacted at 70° C for 16 hours. The mixture was then heated to 90° C for 1 hour to destroy any residual V601, and then was cooled back to 70° C. To the cooled mixture was then added 13.6 g of 95% DBTDL and 41.1 g of TMI. The TMI was added drop wise over the course of approximately 5 minutes while stirring the reaction mixture. Following the procedure of Example 1, the mixture was reacted at 70° C for  
30 approximately 6 hours at which time the reaction was quantitative. The mixture was then cooled to room temperature. The cooled mixture was viscous, transparent solution, containing no visible insoluble mater.

The percent solids of the liquid mixture was determined to be 25.85% using the halogen drying method described above. Subsequent determination of molecular weight was made using the GPC method described above; the copolymer had a  $M_w$  of 251,350 Da and  $M_w/M_n$  of 3.54 based upon two independent measurements. The product was a copolymer of LMA, GMA and HEMA containing random side chains of TMI and was designed herein as LMA/GMA/HEMA-TMI (91/6/3-4.7% w/w) and was suitable for making an organosol containing epoxy reactive groups in the shell.

### Example 5

Using the method and apparatus of Example 1, 2561 g of Norpar<sup>TM</sup> 15, 823 g of LMA, 54 g of 98% HEMA and 8.75 g of V601 were combined and resulting mixture reacted at 70° C for 16 hours. The mixture was then heated to 90° C for 1 hour to destroy any residual V601, and then was cooled back to 70° C. To the cooled mixture was then added 13.6 g of 95% DBTDL and 41.1 g of TMI. The TMI was added drop wise over the course of approximately 5 minutes while stirring the reaction mixture. Following the procedure of Example 1, the mixture was reacted at 70° C for approximately 6 hours at which time the reaction was quantitative. The mixture was then cooled to room temperature. The cooled mixture was viscous, transparent solution, containing no visible insoluble mater.

The percent solids of the liquid mixture was determined to be 25.43% using the halogen drying method described above. Subsequent determination of molecular weight was made using the GPC method described above; the copolymer had a  $M_w$  of 270,765 Da and  $M_w/M_n$  of 3.26 based upon two independent measurements. The product was a copolymer of LMA and HEMA containing random side chains of TMI and was designed herein as LMA/HEMA-TMI (94/6-4.7% w/w) and was suitable for making an organosol containing hydroxy reactive groups in the shell.

The compositions of the graft stabilizers of Example 1 to 5 are summarized in the following Table.

**Table 1. Graft Stabilizers Containing Reactive Groups**

Example Number	Graft Stabilizer Compositions (% w/w)	Solids (%)	Molecular Weight	
			M <sub>w</sub>	M <sub>w</sub> /M <sub>n</sub>
1 (Comparative)	LMA/HEMA-TMI (97/3-4.7)	25.64	231,350	3.24
2	LMA/DAAM/HEMA -TMI (94/3/3-4.7)	24.47	278,800	2.50
3	LMA/MAA/HEMA-TMI (94/3/3-4.7)	25.10	330,300	2.34
4	LMA/GMA/HEMA-TMI (91/6/3-4.7)	25.85	251,350	3.54
5	LMA/HEMA-TMI (94/6-4.7)	25.43	270,765	3.26

## 5 Example 6-13: Preparations of Organosols

### Example 6 (Comparative)

This is a comparative example using the graft stabilizer in Example 1 to prepare an organosol containing non-reactive groups with a core/shell ratio of 8/1. A 5000 ml 3-neck round flask equipped with a condenser, a thermocouple connected to a digital temperature controller, a nitrogen inlet tube connected to a source of dry nitrogen and a magnetic stirrer, was charged with a mixture of 2751 g of Norpar™ 15, 399.9 g of EMA, 97.9 g of BHA, 242.7 g of the graft stabilizer mixture from Example 1 @ 25.64% polymer solids, and 8.40 g of V601. While stirring the mixture, the reaction flask was purged with dry nitrogen for 30 minutes at flow rate of approximately 2 liters/minute. A hollow glass stopper was then inserted into the open end of the condenser and the nitrogen flow rate was reduced to approximately 0.5 liters/minute. The mixture was heated to 70° C for 16 hours. The conversion was quantitative.

Approximately 350 g of n-heptane were added to the cooled organosol, and the resulting mixture was stripped of residual monomer using a rotary evaporator equipped with a dry ice/acetone condenser and operating at a temperature of 90° C and a vacuum of approximately 15 mm Hg. The stripped organosol was cooled to room temperature, yielding an opaque white dispersion.

This organosol was designed LMA/HEMA-TMI//EMA/BHA (97/3-4.7//80/20 %w/w) and can be used to ink formulations which had no function of reactions. The percent solids of the organosol dispersion after stripping was determined to be 15.27% using the halogen drying method described above. Subsequent determination of average particles size was made using the laser diffraction method described above; the organosol had a volume average diameter 32.6  $\mu\text{m}$ .

### Example 7

This example illustrates the use of the graft stabilizer in Example 2 to prepare an organosol containing secondary amine groups in both the core and the shell with a core/shell ratio of 8/1. Using the method and apparatus of Example 6, 2928 g of Norpar<sup>TM</sup> 15, 289.96 g of EMA, 72.49 g of BHA, 10.9 g of DAAM, 190.7 g of the graft stabilizer mixture from Example 2 @ 24.47% polymer solids, and 8.4 g of V601 were combined. The mixture was heated to 70° C for 16 hours. The conversion was quantitative. The mixture then was cooled to room temperature. After stripping the organosol using the method of Example 6 to remove residual monomer, the stripped organosol was cooled to room temperature, yielding an opaque white dispersion. This organosol was designed LMA/DAAM/HEMA-TMI//EMA/DAAM/BHA (94/3/3-4.7//77/3/20 %w/w) and can be used to prepare ink formulations which reacted and formed crosslinked films when fused at high temperature. The fused ink film exhibited improved blocking and erasure resistance. The percent solids of the organosol dispersion after stripping was determined to be 12.15% using the halogen drying method described above. Subsequent determination of average particles size was made using the laser diffraction method described above; the organosol had a volume average diameter of 11.5  $\mu\text{m}$ .

### Example 8

This example illustrates the use of the graft stabilizer in Example 3 to prepare an organosol containing carboxyl groups in both core and shell with a core/shell ratio of 8/1. Using the method and apparatus of Example 6, 2932 g of Norpar<sup>TM</sup> 15, 289.96 g of EMA, 72.49 g of BHA, 10.9 g of MAA, 185.9 g of the graft stabilizer mixture from Example 3



@ 25.10% polymer solids, and 8.4 g of V601 were combined. The mixture was heated to 70° C for 16 hours. The conversion was quantitative. The mixture then was cooled to room temperature. After stripping the organosol using the method of Example 6 to remove residual monomer, the stripped organosol was cooled to room temperature, yielding an opaque white dispersion. This organosol was designed LMA/MAA/HEMA-TMI//EMA/MAA/BHA (94/3/3-4.7//77/3/20 %w/w) and can be used to prepare ink formulations which reacted and formed crosslinked films when fused at high temperature. The fused ink film exhibited improved blocking and erasure resistance. The percent solids of the organosol dispersion after stripping was determined to be 11.31% using the halogen drying method described above. Subsequent determination of average particles size was made using the laser diffraction method described above; the organosol had a volume average diameter of 102.7 µm.

#### Example 9

This example illustrates the use of the graft stabilizer in Example 4 to prepare an organosol containing epoxy groups in both core and shell with a core/shell ratio of 8/1. Using the method and apparatus of Example 6, 2938 g of Norpar<sup>TM</sup> 15, 289.96 g of EMA, 72.49 g of BHA, 10.9 g of GMA, 180.5 g of the graft stabilizer mixture from Example 4 @ 25.85% polymer solids, and 8.4 g of V601 were combined. The mixture was heated to 70° C for 16 hours. The conversion was quantitative. The mixture then was cooled to room temperature. After stripping the organosol using the method of Example 6 to remove residual monomer, the stripped organosol was cooled to room temperature, yielding an opaque white dispersion. This organosol was designed LMA/GMA/HEMA-TMI//EMA/GMA/BHA (91/6/3-4.7//77/3/20 %w/w) and can be used to prepare ink formulations which reacted and formed crosslinked films when fused at high temperature. The fused ink film exhibited improved blocking and erasure resistance. The percent solids of the organosol dispersion after stripping was determined as 11.68% using the halogen drying method described above. Subsequent determination of average particles size was made using the laser diffraction method described above; the organosol had a volume average diameter of 15.5 µm.

### Example 10

This example illustrates the use of the graft stabilizer in Example 5 to prepare an organosol containing hydroxy groups in both the core and shell with a core/shell ratio of 8/1. Using the method and apparatus of Example 6, 2937 g of Norpar<sup>TM</sup> 15, 284.15 g of  
5 EMA, 71.04 g of BHA, 18.1 g of 98% HEMA, 183.5 g of the graft stabilizer mixture from Example 5 @ 25.43% polymer solids, and 6.3g of V601 were combined. The mixture was heated to 70°C for 16 hours. The conversion was quantitative. The mixture then was cooled to room temperature. After stripping the organosol using the method of Example 6 to remove residual monomer, the stripped organosol was cooled to room  
10 temperature, yielding an opaque white dispersion. This organosol was designed LMA/HEMA-TMI//EMA/HEMA/BHA (94/6-4.7//75/5/20 %w/w) and can be used to prepare an ink formulations which reacted and formed crosslinked films when fused at high temperature. The fused ink film exhibited improved blocking and erasure resistance. The percent solids of the organosol dispersion after stripping was determined  
15 to be 11.04% using the halogen drying method described above. Subsequent determination of average particles size was made using the laser diffraction method described above; the organosol had a volume average diameter of 37.9 µm.

### Example 11

This example illustrates the use of the graft stabilizer in Example 1 to prepare an organosol containing isocyanate groups in the core with a core/shell ratio of 8/1. Using the method and apparatus of Example 6, 2349 g of Norpar<sup>TM</sup> 15, 591.99 g of EMA, 148.0 g of BHA, 37.8 g of TMI, 364.9 g of the graft stabilizer mixture from Example 1 @  
20 25.64% polymer solids, and 8.75 g of V601 were combined. The mixture was heated to 70° C for 16 hours. The conversion was quantitative. The mixture then was cooled to room temperature. After stripping the organosol using the method of Example 6 to remove residual monomer, the stripped organosol was cooled to room temperature, yielding an opaque white dispersion. This organosol was designed LMA/HEMA-TMI//EMA/TMI/BHA (97/3-4.7//75/5/20 %w/w) and can be used to prepare ink  
25 formulations which reacted and formed crosslinked films when fused at high temperature. The fused ink film exhibited improved blocking and erasure resistance. The percent  
30

solids of the organosol dispersion after stripping was determined to be 24.54% using the halogen drying method described above. Subsequent determination of average particles size was made using the laser diffraction method described above; the organosol had a volume average diameter of 21.7  $\mu\text{m}$ .

5

### Example 12

This example illustrates the use of the graft stabilizer in Example 3 to prepare an organosol containing carboxyl groups in the shell and isocyanate groups in the core with a core/shell ratio of 8/1. Using the method and apparatus of Example 6, 2322g of Norpar<sup>TM</sup> 15, 591.99 g of EMA, 148.0 g of BHA, 37.8 g of TMI, 364.9 g of the graft stabilizer mixture from Example 3 @ 25.10% polymer solids, and 13.13 g of V601 were combined. The mixture was heated to 70° C for 16 hours. The conversion was quantitative. The mixture then was cooled to room temperature. After stripping the organosol using the method of Example 6 to remove residual monomer, the stripped organosol was cooled to room temperature, yielding an opaque white dispersion. This organosol was designed LMA/MAA/HEMA-TMI//EMA/TMI/BHA (94/3/3-4.7//75/5/20 %w/w) and can be used to prepare ink formulations which reacted and formed crosslinked films when fused at high temperature. The fused ink film exhibited improved blocking and erasure resistance. The percent solids of the organosol dispersion after stripping was determined to be 25.26% using the halogen drying method described above. Subsequent determination of average particles size was made using the laser diffraction method described above; the organosol had a volume average diameter of 5.0  $\mu\text{m}$ .

25

### Example 13

This example illustrates the use of the graft stabilizer in Example 1 to prepare an organosol containing tertiary amine groups in the core with a core/shell ratio of 8/1. Using the method and apparatus of Example 6, 2945 g of Norpar<sup>TM</sup> 15, 362.5 g of EMA, 10.9 g of DMAEMA, 175.2 g of the graft stabilizer mixture from Example 1 @ 25.64% polymer solids, and 6.3 g of V601 were combined. The mixture was heated to 70° C for 16 hours. The conversion was quantitative. The mixture then was cooled to room

30

temperature. After stripping the organosol using the method of Example 6 to remove residual monomer, the stripped organosol was cooled to room temperature, yielding an opaque white dispersion. This organosol was designed LMA/HEMA-TMI/EMA/DMAEMA (97/3-4.7//97/3 %w/w) and can be used to prepare ink

- 5 formulation which reacted and formed crosslinked films when fused at high temperature. The fused ink film exhibited improved blocking and erasure resistance. The percent solids of the organosol dispersion after stripping was determined as 11.67% using the halogen drying method described above. Subsequent determination of average particles size was made using the laser diffraction method described above; the organosol had a
- 10 volume average diameter of 23.7  $\mu\text{m}$ .

The compositions of the organosol copolymers formed in Examples 6-13 are summarized in the following table:

**Table 2. Organosols Containing Reactive Groups**

<b>Example Number</b>	<b>Organosol Compositions (% w/w)</b>	<b>Reactive Group</b>
6 (Comparative)	LMA/HEMA-TMI/EMA-BHA (97/3-4.7//80/20)	None
7	LMA/DAAM/HEMA-TMI/EMA/DAAM/BHA (94/3/3-4.7//77/3/20)	Secondary Amine
8	LMA/MAA/HEMA-TMI/EMA/MAA/BHA (94/3/3-4.7//77/3/20)	Carboxyl
9	LMA/GMA/HEMA-TMI/EMA/GMA/BHA (91/6/3-4.7//77/3/20)	Epoxy
10	LMA/HEMA-TMI/EMA/HEMA/BHA (94/6-4.7//75/5/20)	Hydroxy
11	LMA/HEMA-TMI/EMA/TMI/BHA (97/3-4.7//75/5/20)	Isocyanate
12	LMA/MAA/HEMA-TMI/EMA/TMI/BHA (94/3/3-4.7//75/5/20)	Carboxyl and Isocyanate
13	LMA/HEMA-TMI/EMA/DMAEMA (97/3-4.7//97/3)	Tertiary Amine

#### **Examples 14-18: Preparation of Liquid Toner Compositions.**

For characterization of the prepared liquid toner compositions in these Examples, the following were measured: size-related properties (particle size); charge-related properties (bulk and free phase conductivity, dynamic mobility and zeta potential); and

charge/developed reflectance optical density (Z/ROD), a parameter that is directly proportional to the toner charge/mass (Q/M).

#### **Example 14 (Comparative)**

5 This is a comparative example of preparing a Cyan liquid toner at an organosol pigment ratio of 8.5 using the organosol prepared at a core/shell ratio of 8 in example 6. 279 g of the organosol @ 15.27% (w/w) solids in Norpar<sup>TM</sup>15 were combined with 14 g of Norpar<sup>TM</sup>15, 5 g of Pigment Blue 15:4 (Sun Chemical Company, Cincinnati, Ohio) and 0.90 g of 5.91% Zirconium HEX-CEM solution (OMG Chemical Company, 10 Cleveland, Ohio) in an 8 ounce glass jar. This mixture was then milled in a 0.5 liter vertical bead mill (Model 6TSG-1/4, Amex Co., Ltd., Tokyo, Japan) charged with 390 g of 1.3 mm diameter Potters glass beads (Potters Industries, Inc., Parsippany, NJ). The mill was operated at 2,000 RPM for 1.5 hours without cooling water circulating through the cooling jacket of the milling chamber.

15 A 16% (w/w) solids toner concentrate exhibited the following properties as determined using the test methods described above:

Volume Mean Particle Size: 1.2 micron

Q/M: 324  $\mu\text{C/g}$

Bulk Conductivity: 115 picoMhos/cm

20 Percent Free Phase Conductivity: 10.5%

Dynamic Mobility: 4.46E-12 ( $\text{m}^2/\text{Vsec}$ )

This toner was tested on the printing apparatus described previously. The reflection optical density (OD) was 1.3 at plating voltages greater than 450 volts.

#### **Example 15**

25 This is an example of preparing a Cyan liquid toner which contained epoxy and secondary amine groups that reacted when fused at high temperature. The toner was prepared at organosol pigment ratio 8 by combining the organosols prepared at core/shell ratios of 8 in example 7 and 9. 126 g of the organosol in example 7 @12.15% (w/w) 30 solids in Norpar<sup>TM</sup> 15 and 131 g of the organosol in example 9 @11.68% (w/w) solids in Norpar<sup>TM</sup> 15 were combined with 38 g of Norpar<sup>TM</sup> 15, 4 g of Pigment Blue15:4 (Sun

Chemical Company, Cincinnati, Ohio) and 0.69 g of 5.91% Zirconium HEX-CEM solution (OMG Chemical Company, Cleveland, Ohio) in an 8 ounce glass jar. This mixture was then milled in a 0.5 liter vertical bead mill (Model 6TSG-1/4, Amex Co., Led., Tokyo, Japan) charged with 390 g of 1.3 mm diameter Potters glass beads (Potters Industries, Inc., Parsippany, NJ). The mill was operated at 2,000 RPM for 1.5 hours with cooling water circulating through the cooling jacket of the milling chamber, and the temperature of the chamber was kept at 35° C.

A 12% (w/w) solids toner concentrate exhibited the following properties as determined using the test methods described above:

Volume Mean Particle Size: 4.2 micron  
Q/M: 259  $\mu\text{C/g}$   
Bulk Conductivity: 75 picoMhos/cm  
Percent Free Phase Conductivity: 5.5%  
Dynamic Mobility:  $1.49\text{E-}11$  ( $\text{m}^2/\text{Vsec}$ )

This toner was tested on the printing apparatus described previously. The reflection optical density (OD) was 1.3 at plating voltages greater than 450 volts.

### Example 16

This is an example of preparing a Cyan liquid toner which contained carboxyl and secondary groups that reacted when fused at high temperature. The toner was prepared at organosol pigment ratio 8 by combining the organosols prepared at core/shell ratios of 8 in example 7 and 8. 126 g of the organosol in example 7 @12.15% (w/w) solids in Norpar<sup>TM</sup> 15 and 136 g of the organosol in example 8 @11.31% (w/w) solids in Norpar<sup>TM</sup> 15 were combined with 33 g of Norpar<sup>TM</sup> 15, 4 g of Pigment Blue15:4 (Sun Chemical Company, Cincinnati, Ohio) and 1.73g of 5.91% Zirconium HEX-CEM solution (OMG Chemical Company, Cleveland, Ohio) in an 8 ounce glass jar. This mixture was then milled in a 0.5 liter vertical bead mill (Model 6TSG-1/4, Amex Co., Led., Tokyo, Japan) charged with 390 g of 1.3 mm diameter Potters glass beads (Potters Industries, Inc., Parsippany, NJ). The mill was operated at 2,000 RPM for 1.5 hours with cooling water circulating through the cooling jacket of the milling chamber, and the temperature of the chamber was kept at 35° C.

A 12% (w/w) solids toner concentrate exhibited the following properties as determined using the test methods described above:

Volume Mean Particle Size: 3.6 micron

Q/M: 609  $\mu\text{C/g}$

Bulk Conductivity: 114 picoMhos/cm

Percent Free Phase Conductivity: 5.3%

Dynamic Mobility:  $1.78\text{E-}11$  ( $\text{m}^2/\text{Vsec}$ )

This toner was tested on the printing apparatus described previously. The reflection Optical density (OD) was 1.0 at plating voltages greater than 450 volts.

### Example 17

This is an example of preparing a Cyan liquid toner which contained hydroxy and isocyanate groups that reacted when fused at high temperature. The toner was prepared at organosol pigment ratio 8 by combining the organosols prepared at core/shell ratios of 8 in example 10 and 11. 139 g of the organosol in example 10 @11.04% (w/w) solids in Norpar<sup>TM</sup> 15 and 62 g of the organosol in example 11 @24.54% (w/w) solids in Norpar<sup>TM</sup> 15 were combined with 93 g of Norpar<sup>TM</sup> 15, 4 g of Pigment Blue 15:4 (Sun Chemical Company, Cincinnati, Ohio) and 1.39 g of 5.91% Zirconium HEX-CEM solution (OMG Chemical Company, Cleveland, Ohio) in an 8 ounce glass jar. This mixture was then milled in a 0.5 liter vertical bead mill (Model 6TSG-1/4, Amex Co., Ltd., Tokyo, Japan) charged with 390 g of 1.3 mm diameter Potters glass beads (Potters Industries, Inc., Parsippany, NJ). The mill was operated at 2,000 RPM for 1.5 hours with cooling water circulating through the cooling jacket of the milling chamber, and the temperature of the chamber was kept at 35°C.

A 12% (w/w) solids toner concentrate exhibited the following properties as determined using the test methods described above:

Volume Mean Particle Size: 2.4 micron

Q/M: 664  $\mu\text{C/g}$

Bulk Conductivity: 151 picoMhos/cm

Percent Free Phase Conductivity: 7.0 %

Dynamic Mobility:  $1.67\text{E-}11$  ( $\text{m}^2/\text{Vsec}$ )

This toner was tested on the printing apparatus described previously. The reflection Optical density (OD) was 1.1 at plating voltages greater than 450 volts.

### Example 18

5 This is an example of preparing a Cyan liquid toner which contained carboxyl, isocyanate and hydroxy groups that reacted when fused at high temperature. The toner was prepared at organosol pigment ratio 8 by combining the organosols prepared at core/shell ratios of 8 in example 10 and 12. 139 g of the organosol in example 10 @11.04% (w/w) solids in Norpar™ 15 and 61 g of the organosol in example 12 @25.26% (w/w) solids in Norpar™ 15 were combined with 94 g of Norpar™ 15, 4 g of Pigment Blue 15:4 (Sun Chemical Company, Cincinnati, Ohio) and 1.39 g of 5.91% Zirconium HEX-CEM solution (OMG Chemical Company, Cleveland, Ohio) in an 8 ounce glass jar. This mixture was then milled in a 0.5 liter vertical bead mill (Model 6TSG-1/4, Amex Co., Ltd., Tokyo, Japan) charged with 390 g of 1.3 mm diameter  
15 Potters glass beads (Potters Industries, Inc., Parsippany, NJ). The mill was operated at 2,000 RPM for 1.5 hours with cooling water circulating through the cooling jacket of the milling chamber, and the temperature of the chamber was kept at 35° C.

A 12% (w/w) solids toner concentrate exhibited the following properties as determined using the test methods described above:

20 Volume Mean Particle Size: 3.0 micron  
Q/M: 405  $\mu\text{C/g}$   
Bulk Conductivity: 80 picoMhos/cm  
Percent Free Phase Conductivity: 6.0 %  
Dynamic Mobility:  $1.66\text{E-}11 \text{ (m}^2/\text{Vsec)}$

25 This toner was tested on the printing apparatus described previously. The reflection Optical density (OD) was 1.0 at plating voltages greater than 450 volts.

### Example 19 (Comparative)

30 This is a comparative example of preparing a Cyan liquid toner which only contained epoxy groups and had no function of reaction. The toner was prepared at an organosol pigment ratio of 8 using the organosol prepared at a core/shell ratio of 8 in



example 9. 274 g of the organosol @ 11.68% (w/w) solids in Norpar™15 were combined with 21 g of Norpar™15, 4 g of Pigment Blue 15:4 (Sun Chemical Company, Cincinnati, Ohio) and 0.72g of 5.91% Zirconium HEX-CEM solution (OMG Chemical Company, Cleveland, Ohio) in an 8 ounce glass jar. This mixture was then milled in a  
5 0.5 liter vertical bead mill (Model 6TSG-1/4, Amex Co., Ltd., Tokyo, Japan) charged with 390 g of 1.3 mm diameter Potters glass beads (Potters Industries, Inc., Parsippany, NJ). The mill was operated at 2,000 RPM for 1.5 hours without cooling water circulating through the cooling jacket of the milling chamber.

A 12% (w/w) solids toner concentrate exhibited the following properties as  
10 determined using the test methods described above:

Volume Mean Particle Size: 3.8 micron

Q/M: 235  $\mu\text{C/g}$

Bulk Conductivity: 73 picoMhos/cm

Percent Free Phase Conductivity: 2.9%

15 Dynamic Mobility:  $1.43\text{E-}11 \text{ (m}^2\text{/Vsec)}$

This toner was tested on the printing apparatus described previously. The reflection optical density (OD) was 1.3 at plating voltages greater than 450 volts.

## Example 20

### Erasure Resistance Data

Erasure resistance characteristics of samples were determined to obtain the data in the following table:

5

**Table 3. Liquid Toners**

Example	Reactive Group	Optical Density	Erasure Resistance
14 (Comparative)	None	1.3	Fair
15	Epoxy-Secondary Amine	1.3	Improved
16	Carboxyl-Secondary Amine	1.0	Improved
17	Hydroxy-Isocyanate	1.1	Improved
18	Hydroxy-Carboxyl-Isocyanate	1.0	Improved
19 (Comparative)	Epoxy	1.3	Fair

### Test Method of Image Erasure Resistance

10 The erasure resistance characteristics of the images were tested according to the ASTM F 1319-94. The images were generated on the printing device described previously. The optical density (OD) of the images was kept at 1.3 for cyan, magenta and black and 0.8 for yellow.

The printed images were placed on a Crockmeter (available from Atlas Electric Devices Co., Chicago, IL). A crockmeter test cloth (available from 15 Testfabrics Inc., Middlesex, NJ) was mounted over the end of the finger as suggested by the manufacture. The cloth pass time was recorded and the corresponding OD on the cloth was measured. The erasure resistance of the image was calculated using the following formulation:

20 
$$\text{Erasure Resistance} = \frac{\text{OD}_{(\text{image before erase})} - \text{OD}_{(\text{cloth after erase})}}{\text{OD}_{(\text{image before erase})}}$$

The results are shown in Fig. 4.

Other embodiments of this invention will be apparent to those skilled in the art upon consideration of this specification or from practice of the invention disclosed herein. Various omissions, modifications, and changes to the principles and embodiments described herein may be made by one skilled in the art without departing  
5 from the true scope and spirit of the invention which is indicated by the following claims.

All patents, patent documents, and publications cited herein are hereby incorporated by reference as if individually incorporated.